

Experimental Conditions			
	Reactants		Products
Before reaction	 2 molecules N_2	 3 molecules H_2	0 molecules NH_3
After reaction	 1 molecule N_2	0 molecules H_2	 2 molecules NH_3

- In this reaction, only the hydrogen is completely used up.
- H_2 is the limiting reagent, or the reactant that determines the amount of product that can be formed by a reaction.

12 3 Limiting Reagent And Percent Yield

Deborah Wiegand

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Chemistry in Quantitative Language Christopher O. Oriakhi, 2009 Problem solving is one of the most challenging aspects students encounter in general chemistry courses leading to frustration and failure Consequently many students become less motivated to take additional chemistry courses after their first year This book deals with calculations in general chemistry and its primary goal is to prevent frustration by providing students with innovative, intuitive and systematic strategies to problem solving in chemistry The material addresses this issue by providing several sample problems with carefully explained step-by-step solutions for each concept Key concepts basic theories and equations are provided and worked examples are selected to reflect possible ways problems could be presented to students

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in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Prentice Hall Chemistry,2000 Chemical Principles with Qualitative Analysis William L. Masterton,Emil J. Slowinski,Conrad L. Stanitski,1986

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