

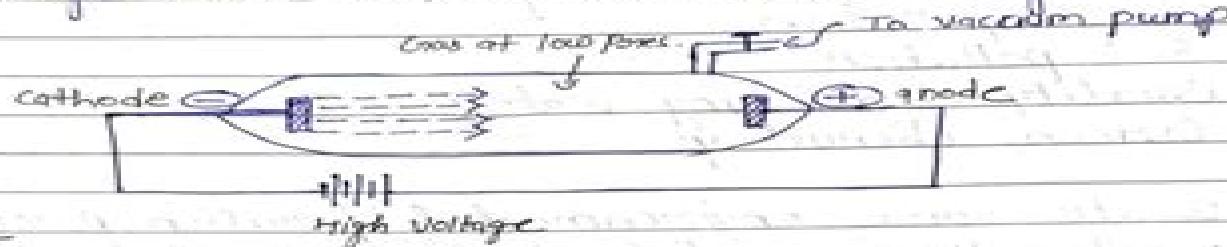
CHAPTER - 2

STRUCTURE OF ATOM

Discovery of electron (Cathode Rays)

Cathode rays \rightarrow William Crookes (1879)
Electron (e^-) \rightarrow J. J. Thomson. (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



Properties

- 1 Cathode rays travel in a straight line (Straight)
- 2 Cathode rays carry -ve charge (negative rays, deflected towards positive)
- 3 Cathode rays are deflected by magnetic field
- 4 They are made up of material particle (Rata, saddle, wool)
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (Other rays)
- 7 They also affect photographic plate.

Discovery of proton (Anode / canal rays)

Anode rays \rightarrow Goldstein (1886)

Proton (p^+) \rightarrow Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

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