

## Review for Ch. 8 Test

## 8.1 Describing Chemical Reactions

- know how to write word equations (i.e. you have be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)
- know how to write a formula equation
- What 2 things do coefficients represent in an equation? **Relative number of moles of a substance ; relative number of molecules needed for a reaction to occur**
- know common symbols used in a reaction
- know 4 evidences of a chemical reaction:
  1. **energy in the form of heat or light**
  2. **production of a gas**
  3. **change in color**
  4. **formation of a precipitate**
- know the Law of conservation of matter: **matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)**
- know what a chemical equation represents: **relative amounts of reactants and products**
- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

1. Complete the equation and write a word equation for the following reaction:  $\text{ZnO} + \text{FeSO}_4 \rightleftharpoons$



**Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide**

2. Write a formula equation for the reaction of a solution of AgCl with potassium metal:



3. What elements are always diatomic gases at room temperature?

**All of the halogens (Group 17) ,  $\text{H}_2$ ,  $\text{N}_2$ ,  $\text{O}_2$**

4. Can you balance these reactions:



# 11 1 Describing Chemical Reactions Practice Problems

## Answer Key

**LM Reder**



## 11 1 Describing Chemical Reactions Practice Problems Answer Key:

**Chemical Interactions** ,2005     *Discover Science: Teacher's annotated edition* ,1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today s technological society     **Soil Chemistry** Daniel G. Strawn,Hinrich L. Bohn,George A. O'Connor,2020-01-28 Provides comprehensive coverage of the chemical interactions among organic and inorganic solids air water microorganisms and the plant roots in soil This book focuses on the species and reaction processes of chemicals in soils with applications to environmental and agricultural issues Topics range from discussion of fundamental chemical processes to review of properties and reactions of chemicals in the environment This new edition contains more examples more illustrations more details of calculations and reorganized material within the chapters including nearly 100 new equations and 51 new figures Each section also ends with an important concepts overview as well as new questions for readers to answer Starting with an introduction to the subject Soil Chemistry 5th Edition offers in depth coverage of properties of elements and molecules characteristics of chemicals in soils soil water chemistry redox reactions in soils mineralogy and weathering processes in soils and chemistry of soil clays The book also provides chapters that examine production and chemistry of soil organic matter surface properties of soil colloids adsorption processes in soils measuring and predicting sorption processes in soils soil acidity and salt affected soils Provides a basic description of important research and fundamental knowledge in the field of soil chemistry Contains more than 200 references provided in figure and table captions and at the end of the chapters Extensively revised with updated figures and tables Soil Chemistry 5th Edition is an excellent text for senior level soil chemistry students     **Prentice Hall Chemistry** ,2000     How to Survive (and Even Excel In) General Chemistry Elizabeth Kean,Catherine Middlecamp,1994 A different kind of book about chemistry which teaches readers the process of learning chemistry not the topic itself Proving a valuable supplement to any introductory text this guide offers inside information to help make chemistry less stressful even enjoyable Includes exercises and sections for self assessment     Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998     Atoms, Molecules, and Reactions Ronald James Gillespie,1994 Traditional college level chemistry including princi     Holt Chemistry Ralph Thomas Myers,2004     **Glencoe Science** McGraw-Hill Staff,2001-06     **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction     General Chemistry with Qualitative Analysis Ralph H. Petrucci,Robert K. Wismer,1987     **ISIJ International** ,2002     **Nieuw Archief Voor Wiskunde** ,1998     **Nuclear Science Abstracts** ,1962     **Understanding Chemistry** Fred M. Dewey,1994     Frank

Schaffer's Chemistry for Everyday David Thurlo,1999 Students learn about important subjects by relating them to events and things that occur in their everyday lives A wealth of interesting activities provide a detailed look into each subject Easy to use activities can be completed individually at school or at home though a few hands on experiments require group work and data sharing A great supplement to any existing curriculum Includes topics such as the scientific method applied to chemistry determining specific gravity balancing chemical equations and exploring the periodic table of elements

*Education Outlook* ,1894      *Educational Times* ,1895      *General Chemistry* Raymond Chang,2003 Publisher

Description      **Perry's Chemical Engineers' Handbook, 9th Edition** Don W. Green,Marylee Z. Southard,2018-07-13 Up to Date Coverage of All Chemical Engineering Topics from the Fundamentals to the State of the Art Now in its 85th Anniversary Edition this industry standard resource has equipped generations of engineers and chemists with vital information data and insights Thoroughly revised to reflect the latest technological advances and processes Perry s Chemical Engineers Handbook Ninth Edition provides unsurpassed coverage of every aspect of chemical engineering You will get comprehensive details on chemical processes reactor modeling biological processes biochemical and membrane separation process and chemical plant safety and much more This fully updated edition covers Unit Conversion Factors and Symbols Physical and Chemical Data including Prediction and Correlation of Physical Properties Mathematics including Differential and Integral Calculus Statistics Optimization Thermodynamics Heat and Mass Transfer Fluid and Particle Dynamics Reaction Kinetics Process Control and Instrumentation Process Economics Transport and Storage of Fluids Heat Transfer Operations and Equipment Psychrometry Evaporative Cooling and Solids Drying Distillation Gas Absorption and Gas Liquid System Design Liquid Liquid Extraction Operations and Equipment Adsorption and Ion Exchange Gas Solid Operations and Equipment Liquid Solid Operations and Equipment Solid Solid Operations and Equipment Chemical Reactors Bio based Reactions and Processing Waste Management including Air Wastewater and Solid Waste Management Process Safety including Inherently Safer Design Energy Resources Conversion and Utilization Materials of Construction

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