

## Stuff I Should Know (Page 2)

### Complex Ions & Common Ligands

Ligands	polar molecules & anions	NH <sub>3</sub> , H <sub>2</sub> O, OH <sup>-</sup> , CN <sup>-</sup> , Cl <sup>-</sup>	Odd example: Fe <sup>2+</sup> + SCN <sup>-</sup> $\rightleftharpoons$ FeSCN <sup>2+</sup>
Central Ions	transition metals and Al <sup>3+</sup>	Ag <sup>+</sup> , Cu <sup>2+</sup> , Ni <sup>2+</sup> , Zn <sup>2+</sup> , etc. & Al <sup>3+</sup>	
Examples	usually twice the number of ligands as the charge on the central ion. Key Words: "excess, concentrated"	Ag(NH <sub>3</sub> ) <sub>2</sub> <sup>+</sup> , Cu(OH) <sub>4</sub> <sup>2-</sup> , Ni(OH) <sub>4</sub> <sup>2-</sup> , Zn(NH <sub>3</sub> ) <sub>6</sub> <sup>2+</sup> , Al(NH <sub>3</sub> ) <sub>6</sub> <sup>3+</sup>	Reaction with Acid: Cu(NH <sub>3</sub> ) <sub>4</sub> <sup>2+</sup> + H <sup>+</sup> $\rightarrow$ Cu <sup>2+</sup> + NH <sub>4</sub> <sup>+</sup>

### Organic Chemistry & Functional Groups

alkenes	alkenes	alkynes	aromatics (benzenes)	nuclear chem	$\Delta H$ , $\Delta S$ , $\Delta G$ ?
C <sub>n</sub> H <sub>2n</sub>	C <sub>n</sub> H <sub>2n</sub>	C <sub>n</sub> H <sub>2n-2</sub>	C <sub>6</sub> H <sub>6</sub>	alpha	= + at all temps
alcohol	aldehyde	ketone	ether	$\frac{1}{2}$ He	= - high temps
R — OH	$\begin{array}{c} \text{O} \\    \\ \text{R} — \text{C} — \text{H} \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{R} — \text{C} — \text{R} \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{R} — \text{O} — \text{R} \end{array}$	beta-electron	= - low temps
carboxylic acid	ester	amine	amide	$\frac{1}{4}$ C	= - no temps
$\begin{array}{c} \text{O} \\    \\ \text{R} — \text{C} — \text{OH} \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{R} — \text{C} — \text{O} — \text{R} \end{array}$	$\begin{array}{c} \text{R} — \text{NH}_2 \\   \\ \text{R} \end{array}$	$\begin{array}{c} \text{O} \\   \\ \text{R} — \text{C} — \text{NH}_2 \end{array}$	neutron	Note: $\Delta S$ in $J$ $\Delta G$ & $\Delta H$ in $kJ$
Substituted benzene	ortho = 1,2	meta = 1,3	para = 1,4	proton	K <sub>a</sub> & Solubility, s.
					1-1 $K_{a,p} = 10^{-1}$
					1-2 $K_{a,p} = 10^{-2}$
					1-3 $K_{a,p} = 10^{-3}$
					2-3 $K_{a,p} = 10^{-10}$

### Lewis Acids & Bases



acid anhydrides (oxides of nonmetals,  $\text{CO}_2$ )

basic anhydrides (oxides of metals,  $\text{MgO}$ )



decomposition reactions:  $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$

Strange Examples:  $\text{P}_2\text{O}_{10} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$

Strange Ions: (outside,  $\text{N}^{3-}$ ) (hydride,  $\text{H}^-$ )



### Flame Test Colors

Barium = green

Sodium = yellow

Copper = blue (w/ green)

Potassium = lavender

Spectron = red

Lithium = red

Calcium = orange

### Quantum Numbers

$n$	1, 2, 3, ...
$l$	0 ... ( $n-1$ )
$m$	-1 ... +1
$m_l$	-0.5 ... +0.5
$\ell$	$0 = s, 1 = p,$ $2 = d, 3 = f$

### Writing Lewis Structures

hint: use one valence electron to connect P's or Cl's then determine lone pairs (Ex:  $\text{K}_3\text{P}_2$ )

### Product-Favored (Spontaneous) Reactions



Properties Indicate Strength of Intermolecular Forces (dMPa)

IMF	BP	FP	H <sub>vap</sub>	H <sub>fus</sub>	VP
IMF	BP	FP	H <sub>vap</sub>	H <sub>fus</sub>	VP

### Orders of Reactions & Graphs That Give Straight Lines

S Order	1 <sup>st</sup> Order	2 <sup>nd</sup> Order
[K] vs. Time	ln[K] vs. Time	$1/[K]$ vs. Time
slope = -k	slope = -k	slope = k

### Electrochemical Cells

anode	cathode
oxidation	reduction
- side	+ side
lower E <sup>o</sup>	higher E <sup>o</sup>
e <sup>-</sup> leave	e <sup>-</sup> enter

### Bond Orders

Bond	B.O.
single	1
double	2
triple	3

### SN & Hydrogenation & Shape

Atomic Number	Hybridization	Basic shape
1	s	
2	sp	linear
3	sp <sup>2</sup>	triangle
4	sp <sup>3</sup>	tetrahedral
5	sp <sup>3</sup> d	bipyramidal
6	sp <sup>3</sup> d <sup>2</sup>	octahedral

### IMF's

London	nonpolar molecules, ex: $\text{CH}_4$ , He
dipole-dipole	polar molecules, ex: $\text{H}_2\text{S}$ , $\text{SO}_2$
hydrogen bonding	$\text{H-F}$ , $\text{H-O}$ , $\text{H-N}$ , $\text{NH}_3$ , $\text{H}_2\text{O}$
metallic	metals, $\text{Ag}$ , $\text{Pb}$
ionic	salts, $\text{NaCl}$ , $\text{CaCO}_3$ , (Note: salts contain covalent bonds)
covalent networks	C(graphite), C(diamond), $\text{SiO}_2$ , $\text{WCl}_6$ , $\text{Si}, \text{SiC}$ (Note: graphite = London, too)

### Activity of Metals (Four Groups)

Metals	React with...
Groups I & II	$\text{H}_2\text{O} \rightarrow \text{Li} + \text{H}_2\text{O} \rightarrow \text{Li}^+ + \text{OH}^- + \text{H}_2$
all others	Non-oxidizing Acid, ex: $\text{HCl}$ $\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$
Cu, Ag, Hg	Oxidizing Acid, $\text{HNO}_3$ or $\text{H}_2\text{SO}_4$ (conc.) $\text{Cu} + \text{HNO}_3 \rightarrow \text{NO}_2 + \text{H}_2\text{O} + \text{Cu}^{2+}$
Al, Pt, Ir	Aqua Regia ( $\text{HNO}_3 + \text{HCl}$ )

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**RC Schank**

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