

**Chemical Equilibrium****STUDY QUESTIONS & PROBLEMS**

- Write the expressions for the equilibrium constant  $K_c$  for the following reactions:
  - $4 \text{NH}_3(\text{g}) + 7 \text{O}_2(\text{g}) \rightleftharpoons 4 \text{NO}_2(\text{g}) + 6 \text{H}_2\text{O}(\text{l})$
  - $\text{HCN}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{CN}^-(\text{aq})$
  - $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
  - $\text{CaCO}_3(\text{s}) \rightleftharpoons \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$
  - $3 \text{O}_2(\text{g}) \rightleftharpoons 2 \text{O}_3(\text{g})$
  - $2 \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq})$
  - $3 \text{Zn}(\text{s}) + 2 \text{Fe}^{3+}(\text{aq}) \rightleftharpoons 2 \text{Fe}(\text{s}) + 3 \text{Zn}^{2+}(\text{aq})$
- Write the equilibrium constant expressions for the following reactions. How are they related to one another?
  - $2 \text{N}_2\text{O}(\text{g}) + 3 \text{O}_2(\text{g}) \rightleftharpoons 4 \text{NO}_2(\text{g})$
  - $\text{N}_2\text{O}(\text{g}) + \frac{3}{2} \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$
  - $4 \text{NO}_2(\text{g}) \rightleftharpoons 2 \text{N}_2\text{O}(\text{g}) + 3 \text{O}_2(\text{g})$
- Calculate the value of the equilibrium constant for the following system, given the data shown:  

$$\text{H}_2(\text{g}) + \text{CO}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g}) + \text{CO}(\text{g})$$

Concentrations at equilibrium:

$[\text{H}_2]$	=	1.5 mol Liter <sup>-1</sup>
$[\text{CO}_2]$	=	2.5 mol Liter <sup>-1</sup>
$[\text{H}_2\text{O}]$	=	0.5 mol Liter <sup>-1</sup>
$[\text{CO}]$	=	3.0 mol Liter <sup>-1</sup>
- Chlorine molecules will dissociate at high temperatures into chlorine atoms. At 3000°C, for example,  $K_c$  for the equilibrium shown is 0.55. If the partial pressure of chlorine molecules is 1.5 atm, calculate the partial pressure of the chlorine atoms:  

$$\text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{Cl}(\text{g})$$
- Suppose that 0.50 moles of hydrogen gas, 0.50 moles of iodine gas, and 0.75 moles of hydrogen iodide gas are introduced into a 2.0 Liter vessel and the system is allowed to reach equilibrium.  

$$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2 \text{HI}(\text{g})$$

Calculate the concentrations of all three substances at equilibrium. At the temperature of the experiment,  $K_c$  equals  $2.0 \times 10^{-2}$ .
- If the mechanism of a chemical equilibrium consists of two reversible elementary steps, each with its own equilibrium constant  $K_{c1}$  and  $K_{c2}$ , what expression relates the equilibrium constant  $K_c$  for the overall equilibrium to the two constants  $K_{c1}$  and  $K_{c2}$ ?
- When 2.0 mol of carbon disulfide and 4.0 mol of chlorine are placed in a 1.0 Liter flask, the following equilibrium system results. At equilibrium, the flask is found to contain 0.20 mol of carbon tetrachloride. What quantities of the other components are present in this equilibrium mixture?  

$$\text{CS}_2(\text{g}) + 3 \text{Cl}_2(\text{g}) \rightleftharpoons \text{S}_2\text{Cl}_2(\text{g}) + \text{CCl}_4(\text{g})$$

# Aqueous Chemical Equilibrium Practice Problems

**Xiaolong Qi**

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