

## Exercises



- Provide an expression for  $K_p$  for reaction 1. Why does a term for C(graphite) not appear in the  $K_p$  expression for the reaction of Table 1?
- Calculate the equilibrium constant for the reaction using tabulated data.

$$\Delta_r H^\circ = -48.4 \text{ kJ/mol}$$

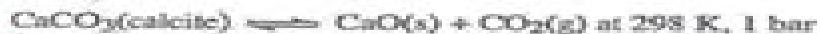


$$\Delta_r S^\circ = -87 \text{ J/K/mol}$$

Write the equilibrium constant for this reaction in terms of the partial pressures.

- Calculate the equilibrium constant for

$$\Delta_r G_f^\circ (\text{CaO}) = -646.3 \text{ kJ/mol}$$



$$\Delta_r G_f^\circ (\text{CO}_2) = -394.36 \text{ kJ/mol}$$

from tabulated data.

$$\Delta_r G_f^\circ (\text{CaO}) = -128.77 \text{ kJ/mol}$$

- The reaction for the formation of nitrosyl chloride



was studied at 25°C. The pressures at equilibrium were found to be

$$P_{\text{NOCl}} = 1.2 \text{ atm}$$

$$P_{\text{NO}} = 5.0 \times 10^{-2} \text{ atm}$$

$$P_{\text{Cl}_2} = 3.0 \times 10^{-1} \text{ atm}$$

- Calculate the value of  $K_p$  for this reaction at 25°C.
- If 1.0 atm of NOCl(g) was placed into an evacuated container would a reaction occur?
- If 1.0 atm each of NOCl(g), Cl<sub>2</sub>(g) and NO(g) were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?
- If 1.0 atm of NOCl(g),  $5.0 \times 10^{-2}$  atm of NO and  $3.0 \times 10^{-1}$  atm of Cl<sub>2</sub> were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

# Ale Chemactivity Answers

**B Lingard**

**Ale Chemactivity Answers:**

The book delves into Ale Chemactivity Answers. Ale Chemactivity Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ale Chemactivity Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Ale Chemactivity Answers
- Chapter 2: Essential Elements of Ale Chemactivity Answers
- Chapter 3: Ale Chemactivity Answers in Everyday Life
- Chapter 4: Ale Chemactivity Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Ale Chemactivity Answers. The first chapter will explore what Ale Chemactivity Answers is, why Ale Chemactivity Answers is vital, and how to effectively learn about Ale Chemactivity Answers.
3. In chapter 2, the author will delve into the foundational concepts of Ale Chemactivity Answers. This chapter will elucidate the essential principles that must be understood to grasp Ale Chemactivity Answers in its entirety.
4. In chapter 3, this book will examine the practical applications of Ale Chemactivity Answers in daily life. This chapter will showcase real-world examples of how Ale Chemactivity Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Ale Chemactivity Answers in specific contexts. This chapter will explore how Ale Chemactivity Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Ale Chemactivity Answers. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Ale Chemactivity Answers.

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