

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b .
3. Calculate the percent ionization of weak acids.

Vocabulary

- | | | |
|----------------------------|-----------------------------|--------------|
| acid dissociation constant | 1. conjugate acid | 2. weak acid |
| acid ionization | 3. base | 4. weak base |
| acid ionization constant | 5. base ionization constant | 6. pK_a |

Key Equations

- | | |
|------------------------------------|-------------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| 3. $pK_a = -\log K_a$ | 4. $pK_b = -\log K_b$ |

Part A Completion

Write the appropriate words in the blanks, and underline the correct word or words. Write the correct word or words in the blanks.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

1. K_a
2. K_b
3. K_a
4. K_b
5. K_a
6. K_b
7. K_a
8. K_b
9. K_a
10. K_b

19 2 Hydrogen Ions And Acidity Section Review

Gary A. Thibodeau, Kevin T. Patton



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