

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the K_a or K_b for the acid or base in parentheses. (For weak acids and bases, include the chemical equation.)

1. HNO_3 (strong acid)
2. $HClO_4$ (strong acid)
3. HCl (strong acid)
4. H_2SO_4 (strong acid)
5. H_2SO_3 (weak acid)
6. H_2CO_3 (weak acid)
7. H_2O (amphoteric)
8. H_2O (amphoteric)
9. H_2O (amphoteric)
10. H_2O (amphoteric)
11. H_2O (amphoteric)
12. H_2O (amphoteric)
13. H_2O (amphoteric)
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20. H_2O (amphoteric)

1. HNO_3
2. $HClO_4$
3. HCl
4. H_2SO_4
5. H_2SO_3
6. H_2CO_3
7. H_2O
8. H_2O
9. H_2O
10. H_2O
11. H_2O
12. H_2O
13. H_2O
14. H_2O
15. H_2O
16. H_2O
17. H_2O
18. H_2O
19. H_2O
20. H_2O

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