

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                       |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution      |
| 4. autoionization | 5. acid rain                | 6. acid rain solution |
|                   | 7. acid solution            | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or pOH. Write the chemical formula for the acid or base. Write the chemical formula for the conjugate acid or base.

1. What is the pH of a 0.1 M solution of hydrochloric acid?
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# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Kathleen Armour**



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the garbage bono thinks troy is cheating on his wife rose

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