

CHEM 12: Hess's Law Worksheet

1. Calculate the ΔH for the reaction: $\text{C}_2\text{H}_6(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$, from the following data:



$$\Delta H = -1411 \text{ kJ / mole}$$
$$\Delta H = -1560 \text{ kJ / mole}$$
$$\Delta H = -285.8 \text{ kJ / mole}$$

2. Calculate the ΔH for the reaction:



from the following data:



$$\Delta H = -180.5 \text{ kJ / mole}$$
$$\Delta H = -91.8 \text{ kJ / mole}$$
$$\Delta H = -483.6 \text{ kJ / mole}$$

3. Find the ΔH_f° (heat of formation) for acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, using the following thermochemical data:



$$\Delta H = -875 \text{ kJ / mole}$$
$$\Delta H = -394.51 \text{ kJ / mole}$$
$$\Delta H = -285.8 \text{ kJ / mole}$$

Pure Factor C(s) is in the form of graphite.

Acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, actually has the structure: CH_3COOH



4. Calculate the ΔH for the reaction: $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + 3\text{H}_2(\text{g})$, from the following data:



$$\Delta H = -91.8 \text{ kJ / mole}$$
$$\Delta H = -74.9 \text{ kJ / mole}$$
$$\Delta H = +270.3 \text{ kJ / mole}$$

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Disha Experts

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