

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b from K_w and K_{eq} for a given chemical reaction involving weak acids and bases.
2. Calculate the pH of a solution of a weak acid or base.

Vocabulary

- acid
- base

- acid-base indicator
- conjugate acid
- conjugate base

- acidic solution
- basic solution
- neutral solution

Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$$

$$K_b = \frac{[\text{B}][\text{H}_3\text{O}^+]}{[\text{BH}^+]}$$

Part A Completion

Write the conjugate acids to the bases and the conjugate bases to the acids. Label each as strong or weak.

1. H_2O
2. H_2O
3. H_2O
4. H_2O
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40. H_2O

19 2 Section Review Hydrogen Ions And Acidity

Linda Swisher, Gary A. Thibodeau



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