

14.3

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pOH | 6. acid solution |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below. Complete each equation as needed.

1. When autoionization is 1.0×10^{-14} , the hydrogen ion concentration is 1.0×10^{-7} . The concentration of hydroxide ion is 1.0×10^{-7} .
2. The hydrogen ion concentration is 1.0×10^{-7} in a 0.1 M solution. The concentration of hydroxide ion is 1.0×10^{-7} .
3. The hydrogen ion concentration is 1.0×10^{-7} in a 0.1 M solution. The concentration of hydroxide ion is 1.0×10^{-7} .
4. The 1.0×10^{-14} autoionization constant is used to find the concentration of hydroxide ion from the concentration of hydrogen ion.
5. The 1.0×10^{-14} autoionization constant is used to find the concentration of hydrogen ion from the concentration of hydroxide ion.

6. 1.0×10^{-14}
7. 1.0×10^{-7}
8. 1.0×10^{-7}
9. 1.0×10^{-7}
10. 1.0×10^{-7}
11. 1.0×10^{-7}
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18. 1.0×10^{-7}

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JR Anderson



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