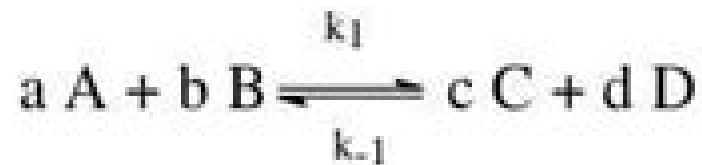


# Chemical Equilibrium



rate of the forward reaction =  $k_1 [A]^a [B]^b$

rate of the reverse reaction =  $k_{-1} [C]^c [D]^d$

At equilibrium:

rate of the forward reaction = rate of the reverse reaction

$$k_1 [A]^a [B]^b = k_{-1} [C]^c [D]^d$$

# 18 Rates Of Reaction Answer

**Karol J. Mysels**



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