

ANSWER KEY: Ch. 17 Acid-base and Solubility Equilibria

Section 17.1

1. **Acid** is a substance that dissociates in water to produce H^+ ions. **Base** is a substance that dissociates in water to produce OH^- ions.

A. The equilibrium system will move away from equilibrium when more acid is added.

B. The equilibrium system will move away from equilibrium when more base is added.

C. The equilibrium system will move away from equilibrium when more acid is added.

D. The equilibrium system will move away from equilibrium when more base is added.

Section 17.2

1. **Acid** is a substance that dissociates in water to produce H^+ ions. **Base** is a substance that dissociates in water to produce OH^- ions.

A. 10^{-1} mol

B. 10^{-2} mol

C. 10^{-3} mol

D. 10^{-4} mol

Section 17.3

1. The equilibrium system will move away from equilibrium when more acid is added.

A. 10^{-1} mol/L

B. 10^{-2} mol/L

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J. A. T. Smellie

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Chemistry Class XI - SBPD Publications Dr. Subhash Chandra Rastogi, , Er. Meera Goyal, 2021-10-16
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Electro Chemistry , Chemical principles Paul B. Kelter, 1998 **Introduction to General, Organic, and Biochemistry** Morris Hein, 1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course. With this in mind, they have anticipated where students might stumble and have paced and organized this text to help them through. Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career. The authors emphasize problem solving and provide a range of practice exercises. As in previous editions, the text first

presents the basic concepts of general chemistry and then moves into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field

General Chemistry Ralph H. Petrucci, William S. Harwood, 1993 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical principles and applications

General Chemistry Raymond Chang, 2003
Publisher Description *The Journal of Physical Chemistry*, 1899 *Maqrin Natural Analogue Study* J. A. T. Smellie, 1998

The Science of Speleology Trevor D. Ford, Cecil Howard Dunstan Cullingford, 1976 **Journal of the Chemical Society** Chemical Society (Great Britain), 1897 Titles of chemical papers in British and foreign journals included in Quarterly journal v

1 12 **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor s resource CD and a solution CD

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