

17 • Acid-Base Equilibria

IN PRACTICE



- Write the pH of each solution above the $[H^+]$ (i.e. $pH = -\log[H^+]$)
- Label the "pH" diagram as "Acidic", "Basic", and "Neutral"
- Remembering that the $[H^+]$ is $[OH^-]$ always equals 1×10^{-14} , fill in the $[OH^-]$ the ends of the two solutions on the "pH" diagram: 1×10^{-14} to reflect the concentration of neutral (i.e. water, $K_w = 1 \times 10^{-14}$ or $10^{-7} \times 10^{-7}$)
- Write the "pOH" of each solution below the $[OH^-]$ (i.e.
- $pH + pOH$ always equals _____
- A solution of acid has $[H^+] = 5.0 \times 10^{-2}$ M
 - Calculate the $[OH^-]$ _____
 - Calculate the pH _____ the pOH _____
- A solution of base has an $[OH^-] = 4.20 \times 10^{-2}$ M
 - Calculate the $[H^+]$ _____
 - Calculate the pH _____ the pOH _____
- Calculate the pH's of the following solutions:
 - 0.01×10^{-1} M HCl pH = _____
 - 0.01×10^{-1} M HCl pH = _____
 - 0.01×10^{-1} M HCl pH = _____
- A pH with 3 significant figures is written with _____ significant after the decimal place.

17 Acid Base Equilibria Ph Practice Answers

John A. Olmsted, Gregory M. Williams



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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