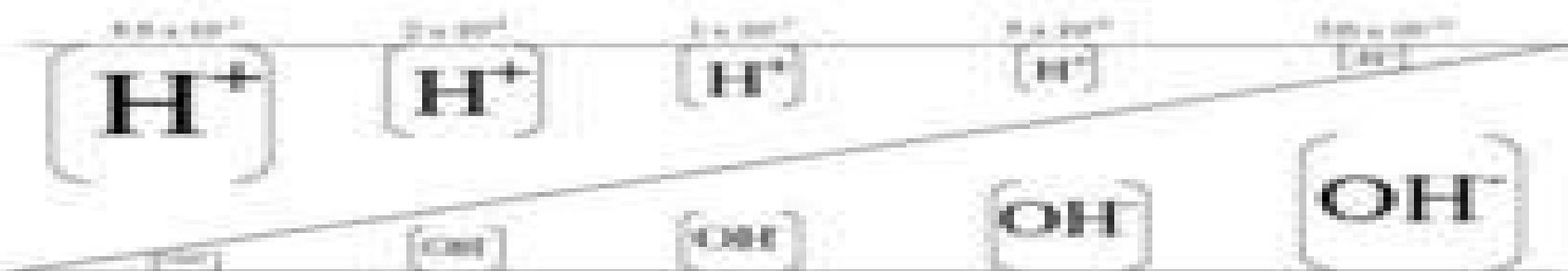


1.7 • Acid-Base Equilibrium

pH = -log₁₀[H⁺]

1. Write the pH of each solution given the following pH or $[\text{H}^+]$:

- 10.00 M H_2SO_4 _____
- 1.00 M H_2SO_4 _____
- 0.001 M H_2SO_4 _____
- 0.0001 M H_2SO_4 _____
- 0.00001 M H_2SO_4 _____
- 0.000001 M H_2SO_4 _____
- 0.0000001 M H_2SO_4 _____
- 0.00000001 M H_2SO_4 _____
- 0.000000001 M H_2SO_4 _____
- 0.0000000001 M H_2SO_4 _____

2. Calculate the pH of aqueous solutions with the following pH or $[\text{H}^+]$:

- $[\text{H}^+] = 10^{-10} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____
- $[\text{H}^+] = 10^{-12} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____
- $[\text{H}^+] = 10^{-14} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____
- $[\text{H}^+] = 10^{-15} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____
- $[\text{H}^+] = 10^{-16} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____
- $[\text{H}^+] = 10^{-17} \text{ M}$
a. Calculate the pH _____
b. Calculate the pOH _____

3. Calculate the pH of a solution with the following pOH :

- $\text{pOH} = 10.00$ _____
- $\text{pOH} = 11.00$ _____
- $\text{pOH} = 12.00$ _____
- $\text{pOH} = 13.00$ _____
- $\text{pOH} = 14.00$ _____

4. Calculate the pH of a solution with the following pH :

- $\text{pH} = 10.00$ _____
- $\text{pH} = 11.00$ _____
- $\text{pH} = 12.00$ _____
- $\text{pH} = 13.00$ _____
- $\text{pH} = 14.00$ _____

5. pH levels of various acids and bases in common acids _____ usually in either the enclosed parentheses.

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John A. Olmsted, Gregory M. Williams

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