

Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius?
a) Li or **K** b) **Ca** or Na c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? from left to right in a period?

Top to bottom: increases due to adding energy levels Left to Right: decreases

3. Why do atoms get smaller as you move left to right in a period?
due to increased nuclear charge and no added shielding

4. Which element in each pair has a larger ionization energy?
a) Na or **O** b) **Be** or Ba c) **Ar** or F d) **Cu** or Ra e) I or **Ne**
f) K or **V** g) **Ca** or Fr h) W or **Se**

5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.
Positive ions are smaller than the neutral atom it comes from, negative ions are larger than the neutral atom it comes from

6. Which particle has the larger radius in each atom/ion pair?
a) Na, Na⁺ b) S, S²⁻ c) I, I⁻ d) **Al**, Al³⁺

7. What is ionization energy? What is first ionization energy?

The amount of energy needed to remove an electron from an atom. First ionization energy is the energy required to remove a single valence electron from an atom.

8. What is the periodic trend for first ionization energy?

It increases going from left to right due to increased nuclear charge, and decreases going down a group due to increased distance from the nucleus and increased shielding.

9. Arrange the following groups of elements in order of increasing ionization energy.
a) Be, Mg, Sr b) Bi, Cs, Ba c) Na, Al, S
high, middle, low high, low, middle low, middle, high

10. Which element in each pair has a higher electronegativity value?
a) Cl, **F** b) C, **N** c) **Mg**, Ne d) As, Ca

11. What is the periodic trend for electronegativity? Increases going left to right, decreases going down.

15 Periodic Trends Answers

Charles H. Corwin



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Table of Contents **15 Periodic Trends Answers**

1. Understanding the eBook **15 Periodic Trends Answers**
 - The Rise of Digital Reading **15 Periodic Trends Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **15 Periodic Trends Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **15 Periodic Trends Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **15 Periodic Trends Answers**
 - Personalized Recommendations
 - **15 Periodic Trends Answers** User Reviews and Ratings
 - **15 Periodic Trends Answers** and Bestseller Lists

5. Accessing 15 Periodic Trends Answers Free and Paid eBooks
 - 15 Periodic Trends Answers Public Domain eBooks
 - 15 Periodic Trends Answers eBook Subscription Services
 - 15 Periodic Trends Answers Budget-Friendly Options
6. Navigating 15 Periodic Trends Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 15 Periodic Trends Answers Compatibility with Devices
 - 15 Periodic Trends Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 15 Periodic Trends Answers
 - Highlighting and Note-Taking 15 Periodic Trends Answers
 - Interactive Elements 15 Periodic Trends Answers
8. Staying Engaged with 15 Periodic Trends Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 15 Periodic Trends Answers
9. Balancing eBooks and Physical Books 15 Periodic Trends Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 15 Periodic Trends Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 15 Periodic Trends Answers
 - Setting Reading Goals 15 Periodic Trends Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 15 Periodic Trends Answers
 - Fact-Checking eBook Content of 15 Periodic Trends Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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