

15-2 Practice Problems

What is the molarity of the solution prepared when 14.3 g of phosphorus trichloride (PCl_3) is dissolved in sufficient water to prepare 2.00 L of solution?

$$\frac{14.3 \text{ g } \text{PCl}_3}{137.33 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0521 \text{ M}$$

Phosphorus pentachloride reacts with water to produce hydrochloric acid. How many grams of phosphorus pentachloride must be added to produce 1.00 L of a 0.10 M solution of phosphorus pentachloride?

$$\frac{1.00 \text{ L} \times 0.10 \text{ M}}{1} \times \frac{1 \text{ mol}}{1} \times \frac{208.2 \text{ g/mol}}{1} = 20.8 \text{ g}$$

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How many grams of phosphorus pentachloride (PCl_5) must be added to produce 1.00 L of a 0.10 M solution of phosphorus pentachloride?

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**Christopher Black, Mark
Anestis, College Hill Coaching
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