

**14-1 Review and Reinforcement (continued)**

Complete each of the sentences below.

15. The stronger the intermolecular forces in a liquid, the \_\_\_\_\_ its boiling point.
16. The state of a substance at room temperature depends on \_\_\_\_\_.
17. Water molecules in an ice cube are held together by \_\_\_\_\_ forces or, more specifically, \_\_\_\_\_ bonds.

On the line at the left, write the letter of the term that matches each description below. Each choice can be used once, more than once, or not at all.

a. dispersion      b. dipole-dipole      c. hydrogen bonding

18. the force in between  $\text{SO}_2$  molecules
19. the force that accounts for HF being a liquid while  $\text{H}_2$  and  $\text{F}_2$  are gases
20. the force that depends on temporary, induced dipoles
21. the only intermolecular force that exists in noble gases
22. the attractive force between neighboring dipoles
23. the force that arises because of the large difference in electronegativities in the N-H bonds

Answer the following question as directed.

24. Which of the following statements could be used to explain why water has an unusually high boiling point? Make a checkmark on the line for all that apply.

- \_\_\_\_\_ a. Water molecules have a strong intermolecular force.
- \_\_\_\_\_ b. There are no nitrogen atoms in water.
- \_\_\_\_\_ c. Oxygen is more electronegative than hydrogen.
- \_\_\_\_\_ d. Water is an angular molecule.
- \_\_\_\_\_ e. Oxygen atoms are smaller than hydrogen atoms.
- \_\_\_\_\_ f. Water does not contain metallic bonds.
- \_\_\_\_\_ g. There is a large difference in electronegativity between oxygen and hydrogen.

## 14 1 Review Reinforcement Answer

**L Cohen**

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