

# Atomic Spectra: Answer Sheet

Name Madi Johnson

Section Number

035

| Substance in the lamp | Color seen with eye | Number of lines observed using glasses (estimate) | Colors of lines in spectrum |
|-----------------------|---------------------|---------------------------------------------------|-----------------------------|
| Hydrogen              | violet              | 3                                                 | Purple, blue, red           |
| Neon                  | red                 | 20                                                | green yellow, orange, red   |
| Mercury               | indigo              | 4                                                 | yellow, green, blue, purple |

| Substance in flame   | Chemical Formula   | Flame Color          |
|----------------------|--------------------|----------------------|
| Lithium chloride     | <chem>LiCl</chem>  | pink, red            |
| Sodium chloride      | <chem>NaCl</chem>  | orange               |
| Potassium chloride   | <chem>KCl</chem>   | blue, orange         |
| Barium chloride      | <chem>BaCl2</chem> | yellow               |
| Copper (II) chloride | <chem>CuCl2</chem> | indigo / green, blue |
| Strontium chloride   | <chem>SrCl2</chem> | red, orange          |

## QUESTIONS

1. A fluorescent light bulb contains hot gaseous mercury atoms. An incandescent bulb contains a thin piece of metal called a filament. It is heated until it glows, a phenomenon known as blackbody radiation. Based on what you know about atomic spectra, which light source will have a continuous spectrum, and which will have a noncontinuous spectrum? Explain.



The fluorescent light bulb will have a non continuous spectrum because as we viewed before mercury had a spectrum broken into lines. The incandescent light bulb will have a continuous spectrum because it is created by thermal energy.

2. Each element has its own distinct emission color bands. Why are the emission spectra of no two elements the same?

The emission spectra of no two elements are the same because each element have a different number of electrons and valence shells.

# Atomic Spectra Lab Answer Key

**International Astronomical Union**

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