

# 7-1 Review and Reinforcement

## Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative	noble gas	octet	electron configuration
positive	electrons	ions	polyatomic
crisscross	anion	cation	empirical formula

- An ionic bond is an attraction between oppositely charged —ions.
- Anions have a —negative charge.
- An atom becomes an ion by losing or gaining electrons.
- The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
- When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
- The atoms that make up polyatomic ions are bonded together by covalent bonds.
- The crisscross method can be used to write the formula for an ionic compound.
- The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

<u>losing</u>	9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons.
<u>true</u>	10. A cation has a positive charge.
<u>true</u>	11. A binary ionic compound contains only one kind of cation and one kind of anion.
<u>polyatomic</u>	12. Monatomic ions consist of more than one atom.
<u>seven</u>	13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.
<u>true (released)</u>	14. A great deal of energy is produced when an ionic compound is formed.

# Atoms And Bonding Chapter 7 Answer Key Prentice Hall

**William Joseph O'Brien**

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*Prentice Hall Chemistry*,2000 **Chemistry** Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 **Prentice Hall Science Explorer** Michael J. Padilla,2002 **Biology**,1998

**Prentice Hall Biology** Sandra Gottfried,Gerry M. Madrazo, Jr.,1990 **Focus on Physical Science California Edition** Michael J. Padilla,2007 **Prentice Hall Science Explorer: Teacher's ed** ,2005 **Inorganic Chemistry** James E. House,2019-11-01 Inorganic Chemistry Third Edition emphasizes fundamental principles including molecular structure acid base chemistry coordination chemistry ligand field theory and solid state chemistry The book is organized into five major themes structure condensed phases solution chemistry main group and coordination compounds each of which is explored with a balance of topics in theoretical and descriptive chemistry Topics covered include the hard soft interaction principle to explain hydrogen bond strengths the strengths of acids and bases and the stability of coordination compounds etc Each chapter opens with narrative introductions and includes figures tables and end of chapter problem sets This new edition features updates throughout with an emphasis on bioinorganic chemistry and a new chapter on nanostructures and graphene In addition more in text worked out examples encourage active learning and prepare students for exams This text is ideal for advanced undergraduate and graduate level students enrolled in the Inorganic Chemistry course Includes physical chemistry to show the relevant principles from bonding theory and thermodynamics Emphasizes the chemical characteristics of main group elements and coordination chemistry Presents chapters that open with narrative introductions figures tables and end

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**Informational Biopolymers of Genes and Gene Expression** R.D. Blake,2005-03-23 This new text examines the biophysics and biochemistry of nucleic acids and proteins carving out the dynamic interface between chemistry and molecular biology and providing a detailed picture of nucleic acids and proteins their structures biological properties and origins and evolution Full of novel insights informed by years of research and teaching R D Blake has written a new text that examines the biophysics and biochemistry of nucleic acids and proteins This book carves out the dynamic interface between chemistry and molecular biology and provides a detailed picture of nucleic acids and proteins their structures biological properties and origins and evolution While pitched at a level that is accessible to upper level undergraduate students of biology the book will likewise be of keen interest to researchers and graduate students in biology chemistry and the physical sciences A broad range of topics accompanied with extensive references footnotes boxed asides chapter ending problems and detailed illustrations throughout combine to make this an ideal choice for courses on advanced biochemistry physical biochemistry nucleic acids and proteins biophysics and biochemistry molecular evolution and protein nucleic acid interactions

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