

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A great deal of energy is <u>produced</u> when an ionic compound is formed. |

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William Joseph O'Brien



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Prentice Hall Chemistry, 2000 **Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 **Prentice Hall Science Explorer** Michael J. Padilla, 2002 *Biology*, 1998 **Prentice Hall Biology** Sandra Gottfried, Gerry M. Madrazo, Jr., 1990 Focus on Physical Science California Edition Michael J. Padilla, 2007 Prentice Hall Science Explorer: Teacher's ed, 2005 *Inorganic Chemistry* James E.

House, 2019-11-01 Inorganic Chemistry Third Edition emphasizes fundamental principles including molecular structure acid base chemistry coordination chemistry ligand field theory and solid state chemistry The book is organized into five major themes structure condensed phases solution chemistry main group and coordination compounds each of which is explored with a balance of topics in theoretical and descriptive chemistry Topics covered include the hard soft interaction principle to explain hydrogen bond strengths the strengths of acids and bases and the stability of coordination compounds etc Each chapter opens with narrative introductions and includes figures tables and end of chapter problem sets This new edition features updates throughout with an emphasis on bioinorganic chemistry and a new chapter on nanostructures and graphene In addition more in text worked out examples encourage active learning and prepare students for exams This text is ideal for advanced undergraduate and graduate level students enrolled in the Inorganic Chemistry course Includes physical chemistry to show the relevant principles from bonding theory and thermodynamics Emphasizes the chemical characteristics of main group elements and coordination chemistry Presents chapters that open with narrative introductions figures tables and end

of chapter problem sets **Inorganic Chemistry** Therald Moeller, 1982-06-18 Presents the entire inorganic field as a logical development of basic ideas incorporating significant early contributions factual data and the resulting modern ideas involving the scope and significance of inorganic chemistry Proceeds sensibly from the origins of the elements through atomic structure molecular structure bonding and properties related to bonding and reactions considered by types conditions and mechanisms **Informational Biopolymers of Genes and Gene Expression** R.D. Blake, 2005-03-23 This new text examines the biophysics and biochemistry of nucleic acids and proteins carving out the dynamic interface between chemistry and molecular biology and providing a detailed picture of nucleic acids and proteins their structures biological properties and origins and evolution Full of novel insights informed by years of research and teaching R D Blake has written a new text that examines the biophysics and biochemistry of nucleic acids and proteins This book carves out the dynamic interface between chemistry and molecular biology and provides a detailed picture of nucleic acids and proteins their structures biological properties and origins and evolution While pitched at a level that is accessible to upper level undergraduate students of biology the book will likewise be of keen interest to researchers and graduate students in biology chemistry and the physical sciences A broad range of topics accompanied with extensive references footnotes boxed asides chapter ending problems and detailed illustrations throughout combine to make this an ideal choice for courses on advanced biochemistry physical biochemistry nucleic acids and proteins biophysics and biochemistry molecular evolution and protein nucleic acid interactions New York Times Saturday Review of Books and Art ,1970 The Science of Zoology Paul B. Weisz, 1973

Supplementary Readings for Chemical Bond Approach Robert K. Fitzgerald, William Franklin Kieffer, 1960 Paperbacks in Print ,1975 **Dental Materials and Their Selection** William Joseph O'Brien, 2002 This is a textbook that includes a general classification of dental materials and guides students and practitioners in the use and selection of these materials

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