

Structure, Bonding and properties



1. Why do atoms bond together?

Some atoms are very reluctant to combine with other atoms and exist in the air around us as single atoms. These are the **Noble Gases** and have **very stable electron arrangements** eg 2, 2.8 and 2.8.8 because their outer shells are full. The first three are shown in the diagrams below and explains why Noble Gases are so reluctant to form compounds with other elements.



helium (2) 2 neon (10) 2.8 Ar (18) 2.8.8 (atomic number) and electron arrangement.

All other atoms therefore, bond together to become electronically more stable, that is to become like Noble Gases in electron arrangement. Bonding produces new substances and usually involves only the 'outer shell' or 'valency' electrons.

The phrase **CHEMICAL BOND** refers to the strong electrical force of attraction between the atoms or ions in the structure. The combining power of an atom is sometimes referred to as its **valency** and its value is linked to the number of outer electrons of the original uncombined atom.

COVALENT BONDING - **sharing electrons** to form molecules with **covalent bonds**, the bond is usually formed between **two non-metallic elements** in a molecule.

IONIC BONDING - By one atom transferring electrons to another atom.

COORDINATE COVALENT BONDING A coordinate covalent bond is special because it involves a shared pair of electrons that came from a single atom.

METALLIC BONDING The crystal lattice of metals consists of ions **NOT atoms** surrounded by a 'sea of electrons' forming giant lattice. These free or 'delocalised' electrons are the 'electronic glue' holding the particles together. There is a **strong electrical force of attraction** between these **mobile electrons (-)** and the 'immobile' **positive metal ions (+)** and this is the **metallic bond**.

- An ion is an atom or group of atoms carrying an overall **positive or negative charge**
 - eg Na^+ , Cl^- , $[\text{Cu}(\text{H}_2\text{O})]^2+$, SO_4^{2-} etc.
- If a particle, as in a neutral atom, has equal numbers of protons (+) and electrons (-) the particle charge is zero ie no overall electric charge.
- The proton number in an atom does not change BUT the number of associated electrons can!
- If negative electrons are removed the excess charge from the protons produces an overall positive ion.
- If negative electrons are gained, there is an excess of negative charge, so a negative ion is formed.
- The charge on the ion is numerically related to the number of electrons transferred.
- For any atom or group of atoms, for every electron gained you get a one unit increase in negative charge, for every electron lost you get a one unit increase in the positive

The atom **losing electrons** forms a **positive ion (cation)** and is usually a **metal**.

The atom **gaining electrons** forms a **negative ion (anion)** and is usually a **non-metallic element**.

Atoms And Bonding Assessment Study Guide

Susan McMurry

Atoms And Bonding Assessment Study Guide:

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