

Name: _____

Date/Period: _____

Atomic Number and Mass Number

Complete the following chart and answer the questions below.

| <i>Element Name</i> | <i>Atomic Number</i> | <i>Number of Protons</i> | <i>Number of Neutrons</i> | <i>Mass Number</i> |
|----------------------------|-----------------------------|---------------------------------|----------------------------------|---------------------------|
| carbon | | | | 12 |
| | 8 | | 8 | |
| hydrogen | | | | 1 |
| | | 2 | | 4 |
| hydrogen | | | 2 | |
| nitrogen | | | | 14 |
| | | | 1 | 2 |
| | 92 | | 146 | |
| Potassium | | | 20 | |
| | 11 | | 12 | |
| | | 47 | | 108 |
| Xenon | | | 77 | |
| | | | 45 | 80 |
| | | 24 | | 52 |
| | | | 16 | 32 |
| Gold | | | | 197 |
| | 76 | | 114 | |

How are the *atomic number* and the *number of protons* related to each other?

How do the *number of protons*, *number of neutrons*, and the *mass number* relate to each other?

What is the *one thing* that determines the identity of an atom?

What sub atomic parts are equal to make a neutral atom?

Explain in your words how to calculate the mass of a sodium atom.

Draw and label three element symbols that are representative of the periodic table. It should have atomic #, mass #, name and element symbol.

Atomic Number Activity Answer

Sir Norman Lockyer



Atomic Number Activity Answer:

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