

# Structure of Atom



## TRY YOURSELF

## ANSWERS

### Topic 1

- Atoms combine in the ratio of small whole numbers to form compounds. In a compound, the relative number and kind of atoms remain constant. This indicates the law of definite proportion which says that in compound the combining elements are present in definite proportion.
- According to Dalton's atomic theory, the elements combine in fixed whole-number ratios to form compounds. Therefore it suggested that compounds are made up of molecules which contain two or more atoms of different elements.
- (i) **Electron** : An electron is a fundamental particle of an atom carrying one unit negative charge and having mass nearly equal to  $\frac{1}{1837}$  of mass of an atom of hydrogen.  
(ii) **Proton** : A proton is a fundamental particle of an atom carrying one unit positive charge and having mass nearly equal to the mass of an atom of hydrogen.  
(iii) **Neutron** : A neutron is a sub-atomic particle carrying no charge and having mass  $1.67 \times 10^{-27}$  kg which is almost equal to that of a hydrogen atom.
- Cathode ray discharge tube experiment by J.J. Thomson showed that electrons are negatively charged particles as on applying an electric field, they get deflected towards positively charged plate.
- Cathode ray discharge tube experiment.



Cathode ray discharge tube experiment

In this experiment an electrical discharge is observed in partially evacuated tube which is made up of glass consisting of two thin pieces of metals called electrodes sealed in it. A very low pressure is maintained and a very high voltage is applied. A stream of particles starts moving from negative electrode (cathode) to the positive electrode (anode) called cathode rays. The behaviour of these rays are similar to negatively charged particles called electrons.

6.

Elements	Protons	Electrons	Neutrons
$^{40}_{18}\text{Ar}$	18	18	22
$^{39}_{19}\text{K}$	19	19	20
$^{23}_{11}\text{Na}$	11	11	12

7.

Species X :  $e = 18$ ,  $p = 15$  and  $n = 16$   
Atomic number ( $Z$ ) =  $p = 15$   
Mass number ( $A$ ) =  $p + n = 15 + 16 = 31$   
Species X is not neutral as  $p \neq e$ . It is an anion with charge equal to excess electrons =  $18 - 15 = 3$ .  
So, symbol is  $^{31}_{15}\text{Cl}^{3-}$ .  
Species Y :  $e = 10$ ,  $p = 13$  and  $n = 14$   
Atomic number ( $Z$ ) =  $p = 13$   
Mass number ( $A$ ) =  $p + n = 13 + 14 = 27$   
Species Y is not neutral as  $p \neq e$ .  
It is a cation with charge equal to deficient electrons =  $13 - 10 = 3$ .

So, symbol is  $^{27}_{13}\text{Al}^{3+}$ .

Species Z :  $e = 36$ ,  $p = 36$  and  $n = 48$   
Atomic number ( $Z$ ) =  $p = 36$   
Mass number ( $A$ ) =  $p + n = 36 + 48 = 84$   
Species Z is neutral as  $p = e$ .  
So, symbol is  $^{84}_{36}\text{Kr}$ .

8. Isoelectronic species are those which have the same number of electrons.

(i)  $\text{Sc}^{3+} = 21 - 3 = 18$  and  $\text{Ti}^{4+} = 22 - 4 = 18$   
This pair is isoelectronic.  
(ii)  $\text{Mn}^{2+} = 25 - 2 = 23$  and  $\text{Fe}^{2+} = 26 - 2 = 24$   
This pair is not isoelectronic.  
(iii)  $\text{Co}^{3+} = 27 - 3 = 24$  and  $\text{Zn}^{2+} = 30 - 2 = 28$   
This pair is not isoelectronic.

9. Work function ( $\phi_0$ ) is the minimum energy required to eject the electron.

$$W_0 = h\nu_0 = \frac{hc}{\lambda_0}$$

$$= \frac{6.626 \times 10^{-34} \times 3 \times 10^8 \text{ m s}^{-1}}{280 \times 10^{-9} \text{ m}} = 7.099 \times 10^{-19} \text{ J}$$

$$10. W_0 = 8.4 \text{ eV}$$

$$= 8.4 \times 1.6 \times 10^{-19} \text{ J}$$

$$= 1.344 \times 10^{-19} \text{ J}$$

( $\because 1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$ )

# Atom Structure Chemthink Answer Key

**Bud Morgan**

## **Atom Structure Chemthink Answer Key:**

**ATOMIC STRUCTURE High School Chemistry Curriculum** Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry. In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure. Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry. We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson, Rutherford and Bohr. Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons, neutrons and electrons. We will also discuss the concept of atomic number, atomic mass and isotopes. In Chapter 3 we will delve into the arrangement of electrons within an atom's electron cloud including energy levels, sublevels and orbital shapes. We will also discuss the principles of electron configuration and how it impacts an atom's chemical properties. Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure. We will explore periodic trends such as atomic radius, ionization energy and electron affinity and discuss how these trends are related to an element's position on the table. In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding, stoichiometry and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science and materials science. Throughout this book we will provide clear explanations, helpful illustrations and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands on learner or a theoretical learner this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

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