

Chapter 2 Practice Problems: Cardinal Numbers and Normality

Source: *Author's calculations*.

- What is the charge on the ion typically formed by each element?
 - oxygen
 - calcium
 - nitrogen
 - aluminum
 - iodine
 - potassium
- How many electrons does the neutral atom gain or lose when each ion forms?
 - Cr^{3+}
 - P^{3-}
 - Li^+
 - Ca^{2+}
 - Cl^-
 - O^{2-}
- Name each ion. Identify each as a cation or an anion.
 - Ba^{2+}
 - Cl^{-}
 - Br^+
 - K^+
 - Na^{2+}
- Write the formula describing the charge's net ionic ion. List under N/A as all necessary.
 - anion
 - cation
 - anion
 - cation
 - anion
 - cation
- Name each of the following ions. Identify each as a cation or an anion.
 - CN^-
 - PO_4^{3-}
 - Ca^{2+}
 - NH_4^+
 - Cl^-
 - NO_3^-

Abstract The following are the abstracts of the papers presented at the 1997 Annual Meeting of the American Psychological Association, held in Washington, DC, from August 1-5, 1997.

- Write the formulas for three binary ionic compounds.
 - magnesium oxide
 - potassium iodide
 - calcium sulfide
 - iron (II) chloride
 - aluminum chloride
 - barium bromide
- Write the formulas for the compounds formed from these pairs of ions.
 - Ba^{2+} and Cl^-
 - Ca^{2+} and O^{2-}
 - Al^{3+} and Cl^-
 - Ag^+ and I^-
 - K^+ and Br^-
 - Fe^{2+} and Cl^-
- Name the following binary ionic compounds.
 - BaCl_2
 - CaCl_2
 - NaCl
 - Fe_2O_3
 - Fe_2O_4
 - BaCl_2
 - Li_2O
 - NaBr
 - K_2O
 - BaCl_2
- Write formulas for the following ionic compounds.
 - calcium phosphate
 - calcium hydroxide
 - potassium chloride
 - magnesium oxide
 - potassium cyanide
 - potassium dichromate
- Write formulas for the following ionic compounds.
 - FeCl_2 and FeCl_3
 - barium ion and hydroxide ion
 - K^+ and Fe^{3+}
 - nitrous ion and carbonate ion

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**Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George**

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