

Kinetics Worksheet 1

3. For the reaction: $A + 2B \rightarrow 3C + D$
the following data were obtained.

run	[A]	[B]	rate M·s ⁻¹
1	0.16	0.36	2.67
2	0.16	0.09	0.67
3	0.32	0.24	7.11
4	0.08	0.09	0.168
5	0.24	0.12	2.0
6	0.02	1.35	?

$$R = k[A]^x[B]^y$$

(A) Write the rate law expression for this reaction.

$$\frac{(0.36)^x}{(0.09)^y} = \frac{2.67 \text{ m s}^{-1}}{0.67 \text{ m s}^{-1}}$$

$$(4)^x = 4$$

$$x \text{ lnd} = \text{lnd} 4$$

$$x = 1$$

$$\frac{(0.16)^y}{(0.08)^y} = \frac{0.67 \text{ m s}^{-1}}{0.168 \text{ m s}^{-1}}$$

$$(2)^y = 4$$

$$y \text{ lnd} 2 = \text{lnd} 4$$

$$y = 2$$

$$R = k[A]^2[B]$$

(B) Calculate the value of k, the rate constant (remember units).

$$k = \frac{R}{A^2 B} = \frac{[0.16]^2 [0.36]}{2.67 \text{ m s}^{-1}} \text{ k} \quad k = \text{m}^{-2} \text{s}^{-1}$$

$$k = 2.90 \text{ m}^{-2} \text{s}^{-1}$$

(C) Calculate the rate of the reaction when [A] = 0.02 M and [B] = 1.35 M.

$$R = 2.90 \text{ m}^{-2} \text{s}^{-1} [0.02]^2 [1.35] = 0.16 \text{ m s}^{-1}$$

(D) Assume that the reaction goes to completion. Under the conditions specified in the second run, what would the final [C] be?



$$0.16A \times \frac{2B}{1A} = 0.32B$$

$$\text{limit} \Rightarrow 0.09B \times \frac{1A}{2B} = 0.045A$$

$$0.16A \times \frac{3C}{1A} = 0.48M$$

$$0.09B \times \frac{3C}{2B} = 0.135M$$

$$0.045A \times \frac{3C}{1A} = \cancel{0.48M} \quad 0.135M C$$

5R

Answers Kinetics Worksheet 1

Michal Rosen-Zvi

Answers Kinetics Worksheet 1:

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