

—Grade 11 chemistry pg. 41 Q1-3 5-6

1. Identify which element in each of the following groups has the largest atomic radius:

alkali metals

2. Arrange the following atoms in order of increasing atomic radius (smallest to largest): K, Rb, Cs, F, Li, C

F > C > Li > K > Rb > Cs

3. Predict which has the larger radius: a magnesium atom or a magnesium ion. Explain.

Magnesium atom would have the larger radius that's because a magnesium ion has more charges so it attracts the nucleus towards it self therefore making it smaller.

5. For each of the following pairs of atoms, state which would have (a) a larger size (b) a higher ionization energy (c) a lower electron affinity

(i) Ca, K (ii) Al, P (iii) I, F (iv) Li, Ra

(i) Ca, K:

- a) k
- b) k
- c) ca

(ii) Al, P:

- a) p
- b) p
- c) all

(iii) I, F:

- a) f
- b) f
- c) i

(iv) Li, Ra:

- a) ra
- b) ra
- c) li

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