

## **ANSWER KEY: BUILD AN ATOM**

### **PART I: ATOM SCREEN**

*Build an Atom simulation (<http://phet.colorado.edu/en/simulation/build-an-atom>)*

1. Explore the *Build an Atom* simulation with your group. As you explore, talk about what you find.
2. a) List two things your group observed in the simulation.  
Responses will vary, but here is what you should see....
  - Adding protons changes the identity of the atom - adding one moves from right to left on the Periodic Table.
  - You can get the atom to be stable/unstable by adjusting the number of neutrons
  - Adding electrons will change it from a +ion to neutral atom, to -ion depending on how many electrons you add
  - Adding electrons - you can only add 2 to the first orbital and 8 to the second
3. b) What particle(s) are found in the center of the atom?  
Protons and neutrons
3. Play until you discover which particle(s) determine(s) the name of the element you build. What did you discover?  
Protons - adding them moves from left to right on the Periodic Table
4. What is the name of the following atoms?
  - a) An atom with 3 protons and 4 neutrons: Lithium
  - b) An atom with 2 protons and 4 neutrons: Helium
  - c) An atom with 4 protons and 4 neutrons: Beryllium
5. Play with the simulation to discover which particles affect the charge of an atom or ion.
  - a) Fill in the blanks below to show your results:

Neutral atoms have the same number of protons and electrons.

Positive ions have more protons than electrons.

Negative ions have fewer protons than electrons.
  - b) Develop a relationship (in the form of a single sentence or equation) that can predict the charge based on the number and types of particle.  
The charge of an atom is determined by the ratio of protons to electrons.

# Answers For Build An Atom Phet

**Graeme Lofts, Merrin J. Evergreen**

## Answers For Build An Atom Phet:

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## **Table of Contents Answers For Build An Atom Phet**

1. Understanding the eBook Answers For Build An Atom Phet
  - The Rise of Digital Reading Answers For Build An Atom Phet
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Build An Atom Phet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers For Build An Atom Phet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Build An Atom Phet
  - Personalized Recommendations
  - Answers For Build An Atom Phet User Reviews and Ratings
  - Answers For Build An Atom Phet and Bestseller Lists
5. Accessing Answers For Build An Atom Phet Free and Paid eBooks
  - Answers For Build An Atom Phet Public Domain eBooks
  - Answers For Build An Atom Phet eBook Subscription Services
  - Answers For Build An Atom Phet Budget-Friendly Options
6. Navigating Answers For Build An Atom Phet eBook Formats

- ePub, PDF, MOBI, and More
- Answers For Build An Atom Phet Compatibility with Devices
- Answers For Build An Atom Phet Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Answers For Build An Atom Phet
- Highlighting and Note-Taking Answers For Build An Atom Phet
- Interactive Elements Answers For Build An Atom Phet

8. Staying Engaged with Answers For Build An Atom Phet

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answers For Build An Atom Phet

9. Balancing eBooks and Physical Books Answers For Build An Atom Phet

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answers For Build An Atom Phet

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answers For Build An Atom Phet

- Setting Reading Goals Answers For Build An Atom Phet
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answers For Build An Atom Phet

- Fact-Checking eBook Content of Answers For Build An Atom Phet
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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