

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|------------------------------------|
| acid | acid dissociation constant (K_a) | acid ionization constant (K_a) |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the K_a expression for the acid dissociation constant for each of the following acids. Write the K_b expression for the base dissociation constant for each of the following bases.

1. HNO_3 (strong acid)
2. H_2SO_4 (strong acid)
3. H_2CO_3 (weak acid)
4. H_2O (amphoteric)
5. H_2O (amphoteric)
6. H_2O (amphoteric)
7. H_2O (amphoteric)
8. H_2O (amphoteric)
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20. H_2O (amphoteric)

1. HNO_3
2. H_2SO_4
3. H_2CO_3
4. H_2O
5. H_2O
6. H_2O
7. H_2O
8. H_2O
9. H_2O
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