

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- k. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- l. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H<sub>2</sub>SO<sub>4</sub>**.

4. **HClO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **small** because  $\text{HClO}_4$  is a strong acid.

c. Dissolve **CaCl<sub>2</sub>** in a portion of water and a portion of vinegar? **common** because it is strong at  $\text{pH} \approx 7$ .

d. In your  $\text{pH}$ -response, is **HClO<sub>4</sub>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate base of **HClO<sub>4</sub>**: **ClO<sub>4</sub><sup>-</sup>**

f. Identify **HClO<sub>4</sub>** as a strong, intermediate, or weak: **strong**, because it loses a proton in the aqueous phase.

5. **H<sub>2</sub>SO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{SO}_4^{2-}$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **small** because it is a strong acid.

c. Dissolve **CaCl<sub>2</sub>** in a portion of water and a portion of vinegar? **common** because it is strong at  $\text{pH} \approx 7$ .

d. In your  $\text{pH}$ -response, is **H<sub>2</sub>SO<sub>4</sub>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate base of **H<sub>2</sub>SO<sub>4</sub>**: **SO<sub>4</sub><sup>2-</sup>**

f. Identify **H<sub>2</sub>SO<sub>4</sub>** as a strong, intermediate, or weak: **strong**

6. **HClO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **small** because **HClO<sub>4</sub>** is a strong acid.

c. Dissolve **CaCl<sub>2</sub>** in a portion of water and a portion of vinegar? **common** because it is strong at  $\text{pH} \approx 7$ .

d. In your  $\text{pH}$ -response, is **HClO<sub>4</sub>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate base of **HClO<sub>4</sub>**: **ClO<sub>4</sub><sup>-</sup>**

f. Identify **HClO<sub>4</sub>** as a strong, intermediate, or weak: **strong**, because it loses a proton in the aqueous phase.

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**KJ Lindholm-Leary**

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