

# Science

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(Chapter – 1) (Matter in Our Surroundings)

(Class – IX)

## Question 4:

Liquids generally have lower density as compared to solids. But you must have observed that ice floats on water. Find out why.

## Answer 4:

The mass per unit volume of a substance is called density (density = mass/volume).

*As the volume of a substance increases, its density decreases.*

Though ice is a solid, it has large number of empty spaces between its particles. These spaces are larger as compared to the spaces present between the particles of water. Thus, the volume of ice is greater than that of water. Hence, the density of ice is less than that of water. A substance with lower density than water can float on water. Therefore, ice floats on water.

## Page 9

## Question 1:

Convert the following temperatures to the celsius scale.

(a) 300 K

(b) 573 K

## Answer 1:

(a)  $300\text{ K} = (300 - 273)^{\circ}\text{C} = 27^{\circ}\text{C}$

(b)  $573\text{ K} = (573 - 273)^{\circ}\text{C} = 300^{\circ}\text{C}$

## Question 2:

Convert the following temperatures to the celsius scale.

(a)  $250^{\circ}$

(b)  $100^{\circ}$

## Answer 2:

(a) Water at  $250^{\circ}\text{C}$  exists in gaseous state.

(b) At  $100^{\circ}\text{C}$ , water can exist in both liquid and gaseous form. At this temperature, after getting the heat equal to the latent heat of vaporization, water starts changing from liquid state to gaseous state.

## Question 3:

For any substance, why does the temperature remain constant during the change of state?

## Answer 3:

During a change of state, the temperature remains constant. This is because all the heat supplied to increase the temperature is utilised (as latent heat) in changing the state by overcoming the forces of attraction between the particles. Therefore, this heat does not contribute in increasing the temperature of the substance.

## Question 4:

Suggest a method to liquefy atmospheric gases.

## Answer 4:

By applying pressure and reducing the temperature, atmospheric gases can be liquefied.

## Page 10

## Question 1:

Why does a desert cooler cool better on a hot dry day?

## Answer 1:

When a liquid evaporates, the particles of the liquid absorb energy from the surroundings to compensate the loss of energy during evaporation. This makes the surroundings cool.

# 2015wacce Science Question

**Mavis Kitcher (Mrs)**



## **2015wacce Science Question:**

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web five of the elements la ce pr nd eu are very reactive and when exposed to air react with oxygen to form an oxide coating that tarnishes the surface for this reason these metals are stored under mineral oil the remainder of the lanthanides are not as reactive and some gd lu retain their silvery metallic appearance for a long time

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web jul 12 2023 lanthanides elements 57 71 are fairly abundant in the earth s crust despite their historic characterization as rare earth elements thulium the rarest naturally occurring lanthanoid is more common in the earth s crust than silver 4 5 10 5 versus 0 79 10 5 by mass

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web may 22 2015 li et al have shown that the biological diversity at lanthanide mining sites decreased when lanthanides were found in high concentrations 4 prompting researchers to isolate microorganisms that can effectively concentrate lanthanides in their cell walls for use in future bioremediation endeavors 5

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web this rooted the idea that the optical properties of ln atoms can be understood in isolation indeed given the interest for the ln iii luminescence most studies focus on understanding the f f transitions

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web jun 30 2023 elements lanthanum 57 through lutetium 71 on the periodic table are considered to be lanthanides lanthanides tend to react with oxygen to form oxides the reaction at room temperature can be slow while heat can cause the reaction to

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web to separate the lanthanides from other elements occurring with them they are chemically combined with specific substances to form lanthanide compounds with low solubility oxalates and fluorides for example a process known as ion

exchange is then used to separate the lanthanides from each other in this process a solution of the

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web dec 12 2019 here s a list of the 15 elements that are lanthanides which run from atomic number 57 lanthanum or ln and 71 lutetium or lu lanthanum symbol ln atomic number 57 cerium symbol ce atomic number 58 praseodymium symbol pr atomic number 59 neodymium symbol nd atomic number 60 promethium symbol pm

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web jan 1 1997 the lanthanides comprise the largest naturally occurring group in the periodic table there are more than 100 minerals known to contain lanthanides but the only two of commercial importance are monazite and bastnaesite the chapter discusses the production of lanthanide elements by electrolysis of fused salts and by metallothermic reduction

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web the fifteen elements shown in table pageindex 1 from lanthanum la 4f 0 to lutetium lu 4f 14 are lanthanoids ln may be used as a general symbol for the lanthanoid elements although lanthanoids scandium sc and yttrium y are sometimes called rare earth elements they are relatively abundant in the earth s crust

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web nov 1 1998 production of lanthanides has shown a continuous increase in recent years 50 taking all these facts into account it is reasonable to consider the development of corrosion protection methods using this family of compounds at first the lanthanide elements principally cerium were used in protection against corrosion at high

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web lanthanides metals peptides and proteins selectivity abstract the essential biological role of rare earth elements lay hidden until the discovery in 2011 that lanthanides are specifically incorporated into a bacterial methanol dehydrogenase

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web jan 1 2000 powell describes displacement chromatography and elution chromatography as the two principal methods for separating lanthanide cations by ion exchange chromatography in displacement chromatography the ions to be separated are first sorbed onto a column of the resin the analyte metals are eluted from the column

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web feb 4 2021 the lanthanides are located on the top row of the two rows of elements found below the main body of the periodic table placing the lanthanides and the actinides the row below the lanthanides below the other elements is just a matter of aesthetics the lanthanides belong to period 6 like cesium barium and hafnium