

# Study Card: Gas Variables

Four variables, referred to as **state variables**, describe every sample of gas. They are:

1. Pressure (**P**)
2. Temperature (**T**)
3. Volume (**V**)
4. Number of moles (**n**)

A change in one of these variables will affect the value of one or more of the other variables.

## Pressure (P)

Gas pressure is caused by gas particles colliding with the container walls and exerting forces upon the walls.

**Pressure = Force / Area**

**Units:** atm, mm Hg, torr, lb/in<sup>2</sup>, N/m<sup>2</sup>

$$1 \text{ atm} = 760 \text{ mm Hg} = 14.7 \text{ lb/in}^2$$

## Temperature (T)

Temperature reflects the average kinetic energy possessed by gas particles as they move about the container.

High temperature gas samples have particles moving at high speeds. Thus, a *thermometer is like a speedometer*.

**Units:** Kelvin (abbreviated K)

## Volume (V)

Volume refers to how much space the gas sample occupies.

**You may know:** Gases are expandable and compressible. Changes in T, P, or n often lead to volume changes.

**Units:** L, mL, m<sup>3</sup>

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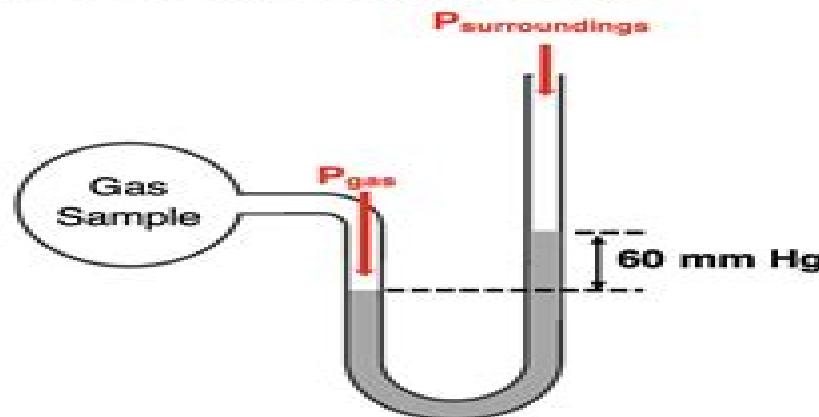
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**Manometers** are used to measure the pressure of a gas. The gas sample presses down on a column of mercury in a U-tube. The difference in height of the Hg levels can be used to determine the gas pressure.



$$P_{\text{gas}} > P_{\text{surroundings}}$$

60 mm height difference

$$\text{If } P_{\text{surroundings}} = 740 \text{ mm Hg}, \\ \text{Then } P_{\text{gas}} = 800 \text{ mm Hg}$$

## Number of Moles (n)

Like pair or dozen, the **mole** is a counting unit used to keep count of the large number of small particles in a sample of matter.

$$1 \text{ mole} = 6.022 \times 10^{23} \text{ particles}$$

We express the number of particles of gas in a sample using the mole. The number of moles can be determined by multiplying the mass of the sample in grams by the molar mass of the gas.

# Gas Variables Packet Answers

**Jicheng Xie**

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