

Electron Configuration (Ions / Valence) - Worksheet #6

Symbol	# of Electrons	Electron Configuration (Use EITHER method. Core or "Long")
Ca^{+2}	18	$\text{Ca}^{+2} : [\text{Ar}]^{18}$
Se^{-2}	34	$\text{Se}^{-2} : [\text{Ar}]^{18} 4s^2 3d^{10} 4p^6$
Sn^{+2}	50	$\text{Sn}^{+2} : [\text{Kr}]^{36} 5s^2 4d^{10}$
Zn^{+2}	28	$\text{Zn}^{+2} : [\text{Ar}]^{18} 3d^{10}$
Sr^{+2}	36	$\text{Sr}^{+2} : [\text{Kr}]^{36}$
O^{-2}	10	$\text{O}^{-2} : [\text{He}]^2 2s^2 2p^6$
Cs^{+1}	54	$\text{Cs}^{+1} : [\text{Xe}]^{54}$
I^{-1}	54	$\text{I}^{-1} : [\text{Xe}]^{54} 5s^2 4d^{10} 5p^6$
Fe^{+3}	23	$\text{Fe}^{+3} : [\text{Ar}]^{18} 4s^1 3d^6$
Cu^{+2}	27	$\text{Cu}^{+2} : [\text{Ar}]^{18} 4s^2 3d^9$

Directions: State how many valence electrons

- O 6
- P 5
- V 5
- Ca 2
- Xe 0
- Hg 0
- Te 6

- Cl 7
- I^{-1} 7
- Zn^{+2} 0
- Ge^{+4} 0
- Sb^{+3} 0
- O^{-2} 0
- Tc^{+4} 0

5 2 Electron Configurations Answer Key

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Chemical Principles, Properties, and Reactions Kenneth L. Henold, Frank Walmsley, 1984

How to Pass Dental Aptitude Tests, Questions and Answers College Publishing Corporation, Brooklyn, 1962

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical

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Table of Contents 5 2 Electron Configurations Answer Key

1. Understanding the eBook 5 2 Electron Configurations Answer Key
 - The Rise of Digital Reading 5 2 Electron Configurations Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 5 2 Electron Configurations Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 5 2 Electron Configurations Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 5 2 Electron Configurations Answer Key
 - Personalized Recommendations
 - 5 2 Electron Configurations Answer Key User Reviews and Ratings
 - 5 2 Electron Configurations Answer Key and Bestseller Lists

5. Accessing 5 2 Electron Configurations Answer Key Free and Paid eBooks
 - 5 2 Electron Configurations Answer Key Public Domain eBooks
 - 5 2 Electron Configurations Answer Key eBook Subscription Services
 - 5 2 Electron Configurations Answer Key Budget-Friendly Options
6. Navigating 5 2 Electron Configurations Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 5 2 Electron Configurations Answer Key Compatibility with Devices
 - 5 2 Electron Configurations Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 5 2 Electron Configurations Answer Key
 - Highlighting and Note-Taking 5 2 Electron Configurations Answer Key
 - Interactive Elements 5 2 Electron Configurations Answer Key
8. Staying Engaged with 5 2 Electron Configurations Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 5 2 Electron Configurations Answer Key
9. Balancing eBooks and Physical Books 5 2 Electron Configurations Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 5 2 Electron Configurations Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 5 2 Electron Configurations Answer Key
 - Setting Reading Goals 5 2 Electron Configurations Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 5 2 Electron Configurations Answer Key
 - Fact-Checking eBook Content of 5 2 Electron Configurations Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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