

## **Chapter 10 Review: Chemical Quantities & Calculations**

*(molar – mole & volume calculations, % composition, empirical & molecular formulas)*

For the test, be sure that you are familiar, and able to complete, the problems below as well as the problems found on the following worksheets: "Molecular Weight and Mole Calculations," "Percent Comp. & Chemical Formula Calculations" and "Chapter 10 Sample Problems." Check the class website for answers to the review worksheet.

1. The air around you is made up of 20.95% oxygen. How many oxygen molecules are in 75.0 grams of air? The density of air is 1.19 g/L. *assume STP.*
2. How much helium (grams) would be needed to inflate a balloon to a volume of 5.50L? *assume STP.*
3. The molecular formula of an antibacterial drug is  $C_{17}H_{18}FN_2O_5$ . How many fluorine atoms are in 150mg of this drug?
4. A molecule is made up of 40.0% carbon, 6.67% hydrogen, and an unknown amount of oxygen. What is the molecular formula of the molecule if its molar mass is 90.0g/mole?
5. What is the empirical formula for a compound made up of 71.72% chlorine, 16.16% oxygen and 12.12% carbon? Calculate the molecular formula for the compound if it has a molar mass of 296.7g/mol.
6. How many molecules are in 8.42 Kilograms of  $C_2H_6$ ?

# 10 Chemical Quantities Guided Practice Problems

## Answers

**Yijin Wang**

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## **Table of Contents 10 Chemical Quantities Guided Practice Problems Answers**

1. Understanding the eBook 10 Chemical Quantities Guided Practice Problems Answers
  - The Rise of Digital Reading 10 Chemical Quantities Guided Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 10 Chemical Quantities Guided Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 10 Chemical Quantities Guided Practice Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 10 Chemical Quantities Guided Practice Problems Answers
  - Personalized Recommendations
  - 10 Chemical Quantities Guided Practice Problems Answers User Reviews and Ratings
  - 10 Chemical Quantities Guided Practice Problems Answers and Bestseller Lists

5. Accessing 10 Chemical Quantities Guided Practice Problems Answers Free and Paid eBooks
  - 10 Chemical Quantities Guided Practice Problems Answers Public Domain eBooks
  - 10 Chemical Quantities Guided Practice Problems Answers eBook Subscription Services
  - 10 Chemical Quantities Guided Practice Problems Answers Budget-Friendly Options
6. Navigating 10 Chemical Quantities Guided Practice Problems Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - 10 Chemical Quantities Guided Practice Problems Answers Compatibility with Devices
  - 10 Chemical Quantities Guided Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 10 Chemical Quantities Guided Practice Problems Answers
  - Highlighting and Note-Taking 10 Chemical Quantities Guided Practice Problems Answers
  - Interactive Elements 10 Chemical Quantities Guided Practice Problems Answers
8. Staying Engaged with 10 Chemical Quantities Guided Practice Problems Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 10 Chemical Quantities Guided Practice Problems Answers
9. Balancing eBooks and Physical Books 10 Chemical Quantities Guided Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 10 Chemical Quantities Guided Practice Problems Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 10 Chemical Quantities Guided Practice Problems Answers
  - Setting Reading Goals 10 Chemical Quantities Guided Practice Problems Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 10 Chemical Quantities Guided Practice Problems Answers
  - Fact-Checking eBook Content of 10 Chemical Quantities Guided Practice Problems Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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