

Multiple Choice:

Select the alternative that best completes the statement or answers the question. What is the most common primary gas used?

1. How many moles of fluorine gas are in a cylinder at 100 atm pressure, 200°C?
- ☐ 1  
☒ 2  
☐ 3  
☐ 4

2. An element has  $2s^2 2p^4$  valence electrons. Which element is it?
- ☒ S  
☐ Se  
☐ O  
☐ Te

3. The compound that does not have a central atom is:
- ☐  $\text{H}_2\text{O}$   
☐  $\text{H}_2\text{O}_2$   
☒  $\text{H}_2\text{SO}_4$   
☐  $\text{H}_2\text{SO}_3$

4. First ionization energy  $\text{Mg} > \text{Al}$  is  $738 \text{ kJ/mol}$  vs.  $578 \text{ kJ/mol}$ . Which of the following best explains this difference?
- ☐ Al has more protons.  
☐ Al has more electrons.  
☒ Al has a smaller atomic radius.  
☐ Al has a higher effective nuclear charge.

5. Which is the most electronegative element?
- ☐ Fluorine  
☒ Oxygen  
☐ Nitrogen  
☐ Carbon

6. Which is the least electronegative element?
- ☐ Fluorine  
☐ Oxygen  
☐ Nitrogen  
☒ Carbon

7. A compound's empirical formula is  $\text{CH}_2$ . At the same time, it has a molar mass of 42 g/mol. What is the molecular formula?
- ☐  $\text{C}_2\text{H}_4$   
☒  $\text{C}_3\text{H}_6$   
☐  $\text{C}_4\text{H}_8$   
☐  $\text{C}_5\text{H}_{10}$

8. First ionization energy  $\text{Mg} > \text{Al}$  is  $738 \text{ kJ/mol}$  vs.  $578 \text{ kJ/mol}$ . Which of the following best explains this difference?
- ☒ Al has more protons.  
☐ Al has more electrons.  
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9. An element has  $2s^2 2p^4$  valence electrons. Which element is it?
- ☐ S  
☐ Se  
☒ O  
☐ Te

10. Which element has the greatest electronegativity?
- ☐ F  
☒ O  
☐ N  
☐ C

11. The electronegativity of N is 3.0. The electronegativity of O is 3.5. The electronegativity of F is 4.0.

# Multiple Choice Stoichiometry Test

**Tao Wei**



## Multiple Choice Stoichiometry Test:

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countries Especially in the case of the school subject chemistry which is very unpopular on the one hand and occupies an exceptional position on the other hand due to its relevance to jobs and everyday life and most notably due to its importance for innovation capacity and problem solving we have to learn from each others educational systems

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