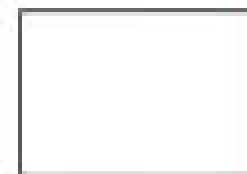
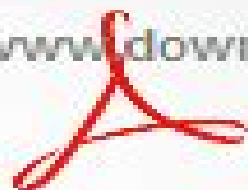


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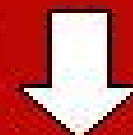


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web determine the final temperature and state of the ammonia once the system reaches equilibrium assuming that no heat is lost to the surroundings use the information below normal boiling point of NH_3 33.3 °C ΔH_{vap} 23.4 kJ/mol specific heat NH_3 4.74 J/g·K specific heat NH_3 2.06 J/g·K

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web apr 3 2023 the experiment described above can be summarized in a graph called a heating curve figure below figure 11.16.1 in the heating curve of water the temperature is shown as heat is continually added changes of state occur during plateaus because the temperature is constant

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web mar 29 2013 heating curves imagine that you have a block of ice that is at a temperature of 30 °C well below its melting point the ice is in a closed container as heat is steadily added to the ice block the water molecules will begin to vibrate faster and faster as they absorb kinetic energy

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web the cooling curve a plot of temperature versus cooling time in figure pageindex 4 plots temperature versus time as a 75 g sample of steam initially at 1 atm and 200 °C is cooled although we might expect the cooling curve to be the mirror image of the heating curve in figure pageindex 3 the cooling curve is not an

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web once all of the material has solidified at the time indicated by point c the cooling will continue at a rate determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase

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we start with a solid and add heat energy for cooling curves we start with the gas phase and remove heat energy cooling and heating curves have five segments let's study each one by looking at heating ice to

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web for water this temperature is 100 °C because the boiling point for water is 100 °C different substances have different melting points and boiling points but the shapes of their heating curves are very similar for example this is the heating curve for iron a metal that melts at 1538 °C and boils at 2861 °C cooling curves

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