

AP BIO: Unit One Study Guide

- Organic compounds: contain covalently bonded carbon; examples include lipids, proteins, nucleic acids and carbohydrates
- Functional groups: amino, carbonyl, carboxyl, hydroxyl, phosphate, and sulfhydryl (*recognize structure*)
- Lipids: non-polymer molecules used for long term energy storage (especially in animals), water-proofing, and insulation
 - o Fat/Oil: glycerol + 3 fatty acids; capital E shape
 - Saturated fat: bad for you; animals and some plants have it; solidifies at room temperature
 - Unsaturated fat: better for you; plants have it; liquefies at room temperature
 - o Steroids: lipids whose structures resemble chicken-wire fence; include cholesterol and sex hormones
 - o Phospholipids: glycerol + 2 fatty acids + 1 phosphate group; make up all membrane bilayers in cells; have hydrophobic interiors (the tails) and hydrophilic exteriors (the heads)
- Carbohydrates: used by cells for energy and structure; monomers are monosaccharides (glucose) or single sugars, disaccharides have two sugars (sucrose, maltose, lactose), and starches are long polymer chains of sugar; storage polysaccharides (starch [plants], glycogen [animals]), structural polysaccharides (chitin [fungal cell walls], cellulose [plant cell walls])
- Proteins: made of amino acid monomers; amino acids consist of an R group attached to an amino and a carboxyl; the amino group of one amino acid can be bonded to the carboxyl group of another amino acid through dehydration synthesis to create a peptide bond and build the protein polymer; proteins serve many functions (e.g., transport, enzymes, cell signals, receptor molecules, structural components, and channels)
 - o Structure: Primary – amino acid polymer made at the ribosome; Secondary – beta pleats or alpha helices; Tertiary – interactions between the R groups of the polymer; Quaternary – interactions with other separate polymer chains (find a buddy)
 - o Enzymes: catalytic proteins that react in an induced-fit fashion with substrates to speed up the rate of reactions by lowering the activation energy; effectiveness is affected by changes in pH, temperature, salinity, etc due to denaturing of structure; examples – pepsin digests proteins and amylase digests starch within the digestive tract
- Nucleic Acids: DNA and RNA; carry the genetic code; monomers are nucleotides consisting of a phosphate group, 5-carbon sugar, and one of five nitrogen bases (adenine, guanine, cytosine, thymine, and uracil)
- ATP: Adenosine triphosphate molecule that acts as a temporary energy storage molecule (battery) for cellular reactions and processes; consisting of an adenosine molecule covalently bonded to three phosphate groups, the addition of the third phosphate group through dehydration synthesis stores small amounts of energy while the removal of the third phosphate group through hydrolysis releases small amounts of energy
- pH: logarithmic scale <7 acidic, 7 neutral, >7 basic (alkaline); pH 4 is 10 times more acidic than pH5; buffers are enzymes that neutralize pH
- Reaction types
 - o Hydrolysis: breaks down polymers by adding water
 - o Dehydration synthesis: creates polymers by removing water
 - o Endergonic reaction: reaction that requires energy input (feels cold); makes bonds
 - o Exergonic reaction: reaction that gives off energy (feels hot); breaks bonds
- Bond types
 - o Covalent bond: sharing electrons; strong
 - o Ionic bond: transferring electrons; strong
 - o Hydrogen bond: temporary attraction between polar molecules; weak; always represented by a dotted line

39 Ap Bio Study Guide Answers

Phillip E. Pack



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Table of Contents 39 Ap Bio Study Guide Answers

1. Understanding the eBook 39 Ap Bio Study Guide Answers
 - The Rise of Digital Reading 39 Ap Bio Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 39 Ap Bio Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 39 Ap Bio Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 39 Ap Bio Study Guide Answers
 - Personalized Recommendations
 - 39 Ap Bio Study Guide Answers User Reviews and Ratings
 - 39 Ap Bio Study Guide Answers and Bestseller Lists

5. Accessing 39 Ap Bio Study Guide Answers Free and Paid eBooks
 - 39 Ap Bio Study Guide Answers Public Domain eBooks
 - 39 Ap Bio Study Guide Answers eBook Subscription Services
 - 39 Ap Bio Study Guide Answers Budget-Friendly Options
6. Navigating 39 Ap Bio Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 39 Ap Bio Study Guide Answers Compatibility with Devices
 - 39 Ap Bio Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 39 Ap Bio Study Guide Answers
 - Highlighting and Note-Taking 39 Ap Bio Study Guide Answers
 - Interactive Elements 39 Ap Bio Study Guide Answers
8. Staying Engaged with 39 Ap Bio Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 39 Ap Bio Study Guide Answers
9. Balancing eBooks and Physical Books 39 Ap Bio Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 39 Ap Bio Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 39 Ap Bio Study Guide Answers
 - Setting Reading Goals 39 Ap Bio Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 39 Ap Bio Study Guide Answers
 - Fact-Checking eBook Content of 39 Ap Bio Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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and Change • Chapter 25 When a radioactive nucleus gives off a gamma ray, its atomic number increases by. 12. The three types of radiation were first identified by Ernest Rutherford. Nuclear Chemistry - Lake Central High School Jul 12, 2015 — What is the change in atomic number after the alpha decay? It decreases by 2. b. ... answer the following questions. Nuclear ... 25.2 Nuclear Transformations | Lecture notes Chemistry These nuclei decay by turning a neutron into a pro- ton to emit a beta particle (an electron) from the nucleus. This process is known as beta emission. It ... 60 s - 1 min SECTION 25.2 NUCLEAR TRANSFORMATIONS. 1. Write a nuclear equation for the following radioactive processes. a. alpha decay of francium-208 $^{208}\text{Fr} \rightarrow$ b ... Cerner Demo 02 PowerChart Basic Overview Part1 - YouTube Basic Cerner training for students - YouTube PowerChart Tutorials | For Medical Professionals eKiDs PowerChart New User Tutorial · Lesson 1: Getting Started · Lesson 2: eKiDs PowerChart Features · Lesson 3: Searching for a Patient · Lesson 4: Opening a ... Cerner General Overview and Structure - YouTube Cerner PowerChart Introduction for Providers - Home Cerner PowerChart Introduction for Providers. Welcome to our Health Quest family! This is a "Flipped Classroom" to get your Cerner PowerChart training started. General Overview of PowerChart - YouTube Cerner Training Bridge Medical Tutorial for Anesthesia Blood Products Transfusion. 3.5K views ... Cerner Radiology Training Series Powerchart Procedure Notes and Autotext Video 3. Cerner Training Video Series Introduction to Order Entry PowerChart Touch Training Open the application to ensure your provider has an access code on his or her device. If you do not have one available, please contact your Cerner Central admin ... PowerChart - Course 205 Building a Patient List. Patient Search. Patient Search Exercise. Banner Bar & Toolbar Functionality. Sticky Note-Question. Sticky Note Exercise.