

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. K_a | 6. K_b |
| | 7. weak solution | |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in 0.00100 mol/L of _____.
3. _____ is the concentration of _____ in 0.100 M _____.
4. _____ is the concentration of _____ in 0.100 M _____.
5. _____ is the concentration of _____ in 0.100 M _____.
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19. _____ is the concentration of _____ in 0.100 M _____.
20. _____ is the concentration of _____ in 0.100 M _____.

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