

# Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



**Reactant side**  
4 Fe atoms  
4 S atoms  
 $7(2) = 14$  O atoms

**Product side**  
4 Fe atoms  
4 S atoms  
 $2(3) + 4(2) = 14$  O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

**Given:**

$m \text{ FeS} = 2.0 \text{ g}$   
**Req'd.** =  $m \text{ O}_2$

**Solution:**

molar mass of  $\text{FeS} = 88 \text{ g/mol}$   
mole ratio = 4 mol  $\text{FeS}$ : 7 mol  $\text{O}_2$   
molar mass of  $\text{O}_2 = 32 \text{ g/mol}$

$$m_{\text{O}_2} = 2.0 \text{ g FeS} \left( \frac{1 \text{ mol FeS}}{88 \text{ g FeS}} \right) \left( \frac{7 \text{ mol O}_2}{4 \text{ mol FeS}} \right) \left( \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) = 1.3 \text{ g O}_2$$

2. Rusting involves the reaction of iron, Fe with oxygen,  $\text{O}_2$  to form iron (III) oxide,  $\text{Fe}_2\text{O}_3$ . (20pts)

a) Write the balanced chemical equation.



**Reactant side**  
4 Fe atoms  
 $3(2) = 6$  O atoms

**Product side**  
4 Fe atoms  
 $2(3) = 6$  O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

**Reactant to each other:**

4 mol Fe: 3 mol  $\text{O}_2$

**Product to each reactant:**

2 mol  $\text{Fe}_2\text{O}_3$ : 4 mol Fe

# Answer Key To Stoichiometry Homework Problems

**Padhraic Smyth**

**Answer Key To Stoichiometry Homework Problems:**

The book delves into Answer Key To Stoichiometry Homework Problems. Answer Key To Stoichiometry Homework Problems is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Answer Key To Stoichiometry Homework Problems, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Answer Key To Stoichiometry Homework Problems
- Chapter 2: Essential Elements of Answer Key To Stoichiometry Homework Problems
- Chapter 3: Answer Key To Stoichiometry Homework Problems in Everyday Life
- Chapter 4: Answer Key To Stoichiometry Homework Problems in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Answer Key To Stoichiometry Homework Problems. This chapter will explore what Answer Key To Stoichiometry Homework Problems is, why Answer Key To Stoichiometry Homework Problems is vital, and how to effectively learn about Answer Key To Stoichiometry Homework Problems.
3. In chapter 2, this book will delve into the foundational concepts of Answer Key To Stoichiometry Homework Problems. The second chapter will elucidate the essential principles that need to be understood to grasp Answer Key To Stoichiometry Homework Problems in its entirety.
4. In chapter 3, this book will examine the practical applications of Answer Key To Stoichiometry Homework Problems in daily life. This chapter will showcase real-world examples of how Answer Key To Stoichiometry Homework Problems can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Answer Key To Stoichiometry Homework Problems in specific contexts. The fourth chapter will explore how Answer Key To Stoichiometry Homework Problems is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Answer Key To Stoichiometry Homework Problems. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Answer Key To Stoichiometry Homework Problems.

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