

Q3 Chapter 9 Preparation (Principles of Chemical Processes and Formulations)

Section 9.10: Processing bases

Q. Which of the changes in the base hydrosol will increase the solubility of the base?

- (a) no change
- (b) addition of water
- (c) addition of more solvents
- (d) evaporation

Q. How many electrons does the neutral form gain or lose before reaching the next ionization state?

- (a) OH^-
- (b) H_2O
- (c) OH_2^+
- (d) OH_2^{2-}
- (e) OH_3^{3-}

Q. Which weak base, similarly weak as ammonia, has the lowest K_b ?

- (a) H_2O
- (b) H_2O_2
- (c) H_2O_3
- (d) H_2O_4
- (e) H_2O_5

Q. Which of the following bases, similarly weak as ammonia, has the lowest K_b ?

- (a) NH_3
- (b) NH_2O
- (c) NH_2O_2
- (d) NH_2O_3
- (e) NH_2O_4

Q.2 Processing and Preparing Formulations that Include OH components

Q. Prepare the following formulary bases (OH components).

- (a) hydrochloric acid
- (b) potassium carbonate
- (c) sodium sulfide
- (d) zinc hydroxide
- (e) aluminum chloride
- (f) lime (calcium hydroxide)

Q. Prepare the following formulary bases (OH components). Identify those bases that possess all three.

- (a) H_2O_2 and H_2O
- (b) H_2O_3 and H_2O
- (c) H_2O_4 and H_2O
- (d) H_2O_5 and H_2O
- (e) H_2O_6 and H_2O

Q. Prepare the following formulary bases (OH components).

- (a) $\text{H}_2\text{O}_2\text{O}_2$
- (b) $\text{H}_2\text{O}_3\text{O}_3$
- (c) $\text{H}_2\text{O}_4\text{O}_4$
- (d) $\text{H}_2\text{O}_5\text{O}_5$
- (e) $\text{H}_2\text{O}_6\text{O}_6$

Q. Prepare formularies for the following water components.

- (a) sodium phosphate
- (b) sodium hydroxide
- (c) ammonium sulfide
- (d) potassium carbonate
- (e) potassium chlorate
- (f) potassium dichromate

Q. Prepare formularies for the following water components.

- (a) $\text{H}_2\text{O}_2\text{O}_2$ and $\text{H}_2\text{O}_3\text{O}_3$
- (b) $\text{H}_2\text{O}_3\text{O}_3$ and $\text{H}_2\text{O}_4\text{O}_4$
- (c) $\text{H}_2\text{O}_4\text{O}_4$ and $\text{H}_2\text{O}_5\text{O}_5$
- (d) $\text{H}_2\text{O}_5\text{O}_5$ and $\text{H}_2\text{O}_6\text{O}_6$
- (e) $\text{H}_2\text{O}_6\text{O}_6$ and $\text{H}_2\text{O}_7\text{O}_7$
- (f) $\text{H}_2\text{O}_7\text{O}_7$ and $\text{H}_2\text{O}_8\text{O}_8$

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