

Chemistry Review Packet

Name: _____

Date: _____

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 16.0 grams B. 26.0 grams C. 34.0 grams D. 44.0 grams E. 50.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A. HNO_3 B. Na_2CO_3 C. HNO_2 D. NaOH E. HClO_2

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Silver B. Baking soda C. Natural gas D. Neon gas E. Water

4. What is the approximate percent composition by mass of hydrogen in the compound ethanol, $\text{C}_2\text{H}_6\text{O}$? (Molecular weight, in grams: C=12.01, H=1.008, O=16.00)

- A. 4.1% B. 12.1% C. 36.5% D. 54.5% E. 20.0%

5. The empirical formula of a certain hydrocarbon is C_2H_5 , and it has a molecular weight equal to 60 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A. C_2H_5 B. C_4H_{10} C. C_6H_{15} D. C_8H_{20} E. $\text{C}_{10}\text{H}_{25}$

6. Which of the following aqueous solutions at 25°C has the highest normal boiling point?

- A. 0.20 M CaCl_2 B. 0.20 M NaCl C. 0.5 M KNO_3 D. 1.0 M $\text{C}_6\text{H}_{12}\text{O}_6$ E. 0.20 M MgCl_2

7. Consider the reaction



How is this reaction best classified?

- A. Precipitation B. Redox reaction C. Single replacement D. Combustion E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 10.0 L of hydrogen gas reacts completely with excess nitrogen, how many liters of ammonia form?

- A. 1.00 L B. 1.50 L C. 40.0 L D. 10.0 L E. 20.0 L

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Table of Contents Answers Chemistry Review Packet

1. Understanding the eBook Answers Chemistry Review Packet
 - The Rise of Digital Reading Answers Chemistry Review Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Chemistry Review Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Chemistry Review Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Chemistry Review Packet

- Personalized Recommendations
- Answers Chemistry Review Packet User Reviews and Ratings
- Answers Chemistry Review Packet and Bestseller Lists
- 5. Accessing Answers Chemistry Review Packet Free and Paid eBooks
 - Answers Chemistry Review Packet Public Domain eBooks
 - Answers Chemistry Review Packet eBook Subscription Services
 - Answers Chemistry Review Packet Budget-Friendly Options
- 6. Navigating Answers Chemistry Review Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Chemistry Review Packet Compatibility with Devices
 - Answers Chemistry Review Packet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Chemistry Review Packet
 - Highlighting and Note-Taking Answers Chemistry Review Packet
 - Interactive Elements Answers Chemistry Review Packet
- 8. Staying Engaged with Answers Chemistry Review Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Chemistry Review Packet
- 9. Balancing eBooks and Physical Books Answers Chemistry Review Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Chemistry Review Packet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Chemistry Review Packet
 - Setting Reading Goals Answers Chemistry Review Packet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Chemistry Review Packet

- Fact-Checking eBook Content of Answers Chemistry Review Packet
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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


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