

IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to obtain a stable octet of electrons at least part of the time. For example:

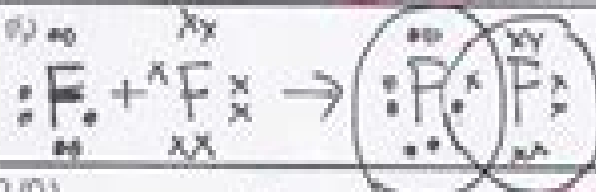


Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO})$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



Answer Key To The Nature Of Covalent Bonding

G Thomas



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Table of Contents Answer Key To The Nature Of Covalent Bonding

1. Understanding the eBook Answer Key To The Nature Of Covalent Bonding
 - The Rise of Digital Reading Answer Key To The Nature Of Covalent Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To The Nature Of Covalent Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To The Nature Of Covalent Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To The Nature Of Covalent Bonding
 - Personalized Recommendations
 - Answer Key To The Nature Of Covalent Bonding User Reviews and Ratings
 - Answer Key To The Nature Of Covalent Bonding and Bestseller Lists
5. Accessing Answer Key To The Nature Of Covalent Bonding Free and Paid eBooks
 - Answer Key To The Nature Of Covalent Bonding Public Domain eBooks
 - Answer Key To The Nature Of Covalent Bonding eBook Subscription Services

- Answer Key To The Nature Of Covalent Bonding Budget-Friendly Options
- 6. Navigating Answer Key To The Nature Of Covalent Bonding eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To The Nature Of Covalent Bonding Compatibility with Devices
 - Answer Key To The Nature Of Covalent Bonding Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To The Nature Of Covalent Bonding
 - Highlighting and Note-Taking Answer Key To The Nature Of Covalent Bonding
 - Interactive Elements Answer Key To The Nature Of Covalent Bonding
- 8. Staying Engaged with Answer Key To The Nature Of Covalent Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To The Nature Of Covalent Bonding
- 9. Balancing eBooks and Physical Books Answer Key To The Nature Of Covalent Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To The Nature Of Covalent Bonding
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key To The Nature Of Covalent Bonding
 - Setting Reading Goals Answer Key To The Nature Of Covalent Bonding
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key To The Nature Of Covalent Bonding
 - Fact-Checking eBook Content of Answer Key To The Nature Of Covalent Bonding
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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