

Periodic Trends

ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
a) Al B b) S O c) Br Cl
d) Na Al e) O F f) Mg Ca
6. Put the following elements in order from smallest to largest atomic radius **and explain why**:
C, O, Sn, Sr.

ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
a) Ca Ga b) Li O c) Cl S d) Br As e) Ba Sr f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions? _____
15. Which group tends to form +2 ions? _____
16. Which group tends to form -1 ions? _____
17. Which group tends not to form ions or react? _____

Atomic Structure Graphing Periodic Trends Answer Key

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Table of Contents Atomic Structure Graphing Periodic Trends Answer Key

1. Understanding the eBook Atomic Structure Graphing Periodic Trends Answer Key
 - The Rise of Digital Reading Atomic Structure Graphing Periodic Trends Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Graphing Periodic Trends Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Graphing Periodic Trends Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Graphing Periodic Trends Answer Key
 - Personalized Recommendations
 - Atomic Structure Graphing Periodic Trends Answer Key User Reviews and Ratings
 - Atomic Structure Graphing Periodic Trends Answer Key and Bestseller Lists

5. Accessing Atomic Structure Graphing Periodic Trends Answer Key Free and Paid eBooks
 - Atomic Structure Graphing Periodic Trends Answer Key Public Domain eBooks
 - Atomic Structure Graphing Periodic Trends Answer Key eBook Subscription Services
 - Atomic Structure Graphing Periodic Trends Answer Key Budget-Friendly Options
6. Navigating Atomic Structure Graphing Periodic Trends Answer Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - Atomic Structure Graphing Periodic Trends Answer Key Compatibility with Devices
 - Atomic Structure Graphing Periodic Trends Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atomic Structure Graphing Periodic Trends Answer Key
 - Highlighting and Note-Taking Atomic Structure Graphing Periodic Trends Answer Key
 - Interactive Elements Atomic Structure Graphing Periodic Trends Answer Key
8. Staying Engaged with Atomic Structure Graphing Periodic Trends Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Graphing Periodic Trends Answer Key
9. Balancing eBooks and Physical Books Atomic Structure Graphing Periodic Trends Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atomic Structure Graphing Periodic Trends Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Atomic Structure Graphing Periodic Trends Answer Key
 - Setting Reading Goals Atomic Structure Graphing Periodic Trends Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atomic Structure Graphing Periodic Trends Answer Key
 - Fact-Checking eBook Content of Atomic Structure Graphing Periodic Trends Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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