

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly empty space with tiny particles.
- Charged particles move in a small elliptical path around a central nucleus.
- This can only hold the negatively charged particles in the center.

BOHR'S MODEL

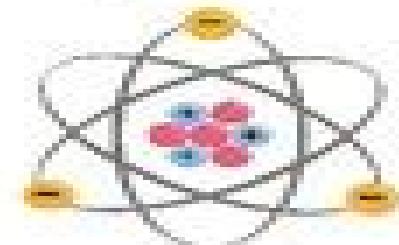
- Shows the elliptical path of an electron.
- Electrons
- Electrons

THE JELLY PUDDING MODEL



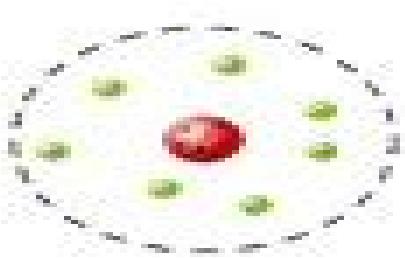
WAVE-MECHANICAL MODEL

- Electrons move in elliptical paths around the nucleus.
- Electrons move in elliptical paths around the nucleus.
- Electrons move in elliptical paths around the nucleus.



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals.
 - Electrons DO move around the nucleus in different patterns.



| Subatomic Particle | Location in the Atom | Charge | Weight |
|--------------------|---|---------------------|--|
| proton | center (in the nucleus) | positive | 1.67 x 10 ⁻²⁴ |
| neutron | center (in the nucleus) | neutral (no charge) | 1.67 x 10 ⁻²⁴ |
| electron | around the nucleus in elliptical orbits | negative | 1.67 x 10 ⁻²⁴ x 10 ⁻³⁰ |

IONS

- If an atom gains or loses an electron, it becomes an ion.
- This gives the atom a charge:
 - Loss of electrons, the atom becomes positive.
 - Gains of electrons, the atom becomes negative.

ISOTOPES

- An isotope is an atom that has a different mass number.
- An isotope is still the same element, it is just a little different from every other atom of that same element.

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