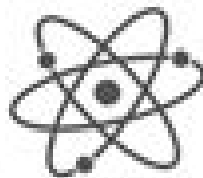


ATOMIC STRUCTURE

Atomic Structure

Atoms are the building blocks of all matter. They are incredibly small, and they are made up of even smaller particles called protons, neutrons, and electrons. These particles determine the properties and behavior of an atom.



Protons are positively charged particles that are found in the **nucleus** (center) of an atom. The number of protons in an atom is called the **atomic number**, and it determines what **element** the atom is. For example, all hydrogen atoms have one proton in their nucleus, while all helium atoms have two protons.

Neutrons are neutral particles that also reside in the nucleus of an atom. They help to hold the protons in place, but they do not contribute to the atomic number. The number of neutrons in an atom can vary. Atoms with the same number of protons but a different number of neutrons are called **isotopes**.

Electrons are negatively charged particles that orbit the nucleus in shells called **energy levels**. The number of electrons in an atom is always equal to the number of protons, but they are arranged in a specific way. The outermost shell of an atom is called the **valence shell**, and it determines how an atom will interact with other atoms to form chemical bonds.

Atoms can be combined to form **molecules**, which in turn make up everything in the world around us, from the air we breathe to the food we eat. The way atoms interact with each other is determined by the arrangement of their electrons, and this is known as **chemical bonding**.

By studying atomic structure, scientists can learn about the properties of different elements and how they interact with each other. This knowledge can be applied in many fields, including medicine, *astronomy and materials science. For example, understanding how atoms*

Atomic Structure

Name: _____	Date: _____
Q.1 What are atoms made up of?	Q.2 What is the name of the positively charged particle found in the nucleus of an atom?
a. Only Electrons	a. Electron
b. Protons and neutrons	b. Neutron
c. Electrons and protons	c. Proton
d. Electrons, protons, and neutrons	d. Nucleon
Q.3 What is the atomic number?	Q.4 Isotopes have same number of:
a. Number of protons in an atom	a. neutrons
b. Number of neutrons in an atom	b. nucleons
c. Number of electrons in an atom	c. protons
d. Sum of protons and neutrons in an atom	d. all of the above
Q.5 Which of the following is a negatively charged particle?	Q.6 What is the outermost shell of an atom called?
a. Electron	a. The valence shell
b. Neutron	b. The electron shell
c. Proton	c. The energy level
d. None of the above	d. The proton shell
Q.7 What is a molecule?	Q.8 How many protons are in a helium atom?
a. A group of atoms bonded together	a. 1
b. A single atom	b. 2
c. A type of chemical bond	c. 3
d. A mixture of elements	d. 4
Q.9 What is an example of a molecule?	
a. Water	

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Carol C. Kanar, Elizabeth Grasby



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