

16. Increasing the temperature generally ...

- A. decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
- B. increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
- C. increases the solubility of both a solid and a gas in a liquid.
- D. decreases the solubility of both a solid and a gas in a liquid.

17. A white precipitate was formed when $\text{BaCl}_{2(\text{aq})}$ was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid effervescence. Salt X is likely to contain

- A. SO_4^{2-} ions.
- B. SO_3^{2-} ions.
- C. NO_3^- ions.
- D. CO_3^{2-} ions.

18. Before a reaction could take place, there should be

- A. ionization of reactant particles.
- B. breakage of bonds of products.
- C. breakage of bonds of reactants.
- D. ionization of product particles.

19. Consider the following reaction equation:
$$\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$$

Silicon(IV) oxide is acting as

- A. a basic oxide.
- B. a reducing agent.
- C. an acidic oxide.
- D. an oxidizing agent.

20. Which of the following acids would form normal salt only?

- A. Tetraoxosulphate(VI) acid
- B. Trioxonitrate(V) acid
- C. Tetraoxophosphate(V) acid
- D. Trioxosulphate(IV) acid

21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm³ of oxygen and 900 cm³ of nitrogen gas?

- A. 1.0 atm
- B. 0.1 atm
- C. 9.0 atm
- D. 0.9 atm

22. Graphite is used as a dry lubricant due to the presence of

- A. mobile electrons.
- B. free electrons.
- C. octahedral structures.
- D. layered structures.

23. Which of the following gases is alkaline?

- A. NO_2
- B. CO_2
- C. NH_3
- D. N_2O

24. Which of the following statements about an equilibrium system is correct?

- A. Forward and backward reactions occur at the same rate.
- B. The concentrations of reactants must equal that of the products.
- C. The concentrations of reactants and products can be changed by adding a catalyst.
- D. The concentrations of reactants and products are not affected by a change in temperature.

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