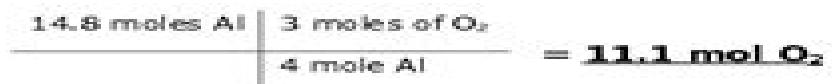


## Worksheet #3—Stoichiometry

**Directions:** Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation:  $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$



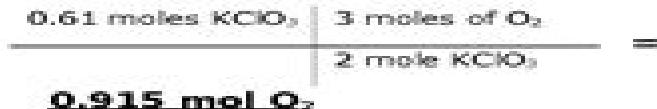
3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride ( $\text{ZnCl}_2$ ) are produced?



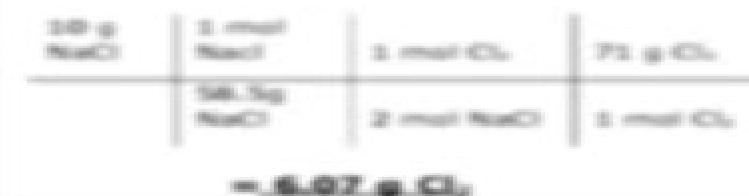
5. How many grams of copper are required to precipitate 4.00 grams of silver nitrate, which were dissolved in reaction?



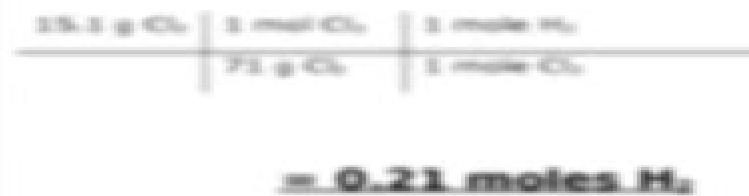
2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate ( $\text{KClO}_3$ ).



4. How many grams of chlorine gas must be reacted with excess sodium iodide ( $\text{NaI}$ ) if 10.0 grams of sodium chloride ( $\text{NaCl}$ ) are needed?



6. How many moles of hydrogen gas are required to react with 15.3 g chlorine gas to produce hydrogen chloride gas?



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**Becker,Charles F. Batten**

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