

Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that 1 mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52°C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86°C, to -1.86°C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pure solvent. $\Delta T_b = i K_b m$ ΔT_b = change in boiling temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_b = molal boiling point elevation constant (°C/m) For H₂O: K_b = .52°C/m <i>m</i> = molality of the solution	Freezing Point Depression: A solution will freeze at a lower temperature than the pure solvent. $\Delta T_f = i K_f m$ ΔT_f = change in freezing temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_f = molal freezing point lowering constant (°C/m) For H₂O: K_f = 1.86°C/m <i>m</i> = molality of the solution
	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_b .	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_f .

EXAMPLE #1

If 90.0 g of nonionizing C₆H₁₂O₆, are dissolved in 255 g of H₂O, what will be the boiling point of the resulting solution?

First calculate the molality (*m*) of the solution:

$$m = \frac{92\text{g}}{\left(180 \frac{\text{g}}{\text{mole}}\right)(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation:

$$[\textit{i} = 1 \text{ for a nonionizing solute.}]$$

$$\Delta T_b = i K_b m$$

$$\Delta T_b = (1)(0.52^\circ\text{C/m})(2.0 \text{ m})$$

$$\Delta T_b = 1.04^\circ\text{C}$$

$$\text{BP} = (\Delta T_b + 100.00^\circ\text{C})$$

$$\text{BP} = (1.04^\circ\text{C} + 100.00^\circ\text{C}) = \underline{101.04^\circ\text{C}}$$

16 3 Colligative Properties Of Solutions Section Review

Stephen D. Bresnick



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application Foundations of College Chemistry Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity **5 Steps to a 5: AP Chemistry 2021** Mary Millhollon, Richard H. Langley, 2020-10-02 MATCHES THE LATEST EXAM In this hybrid year let us supplement your AP classroom experience with this multi platform study guide The immensely popular 5 Steps to a 5 AP Chemistry guide has been updated for the 2020 21 school year and now contains 3 full length practice exams available both in the book and online that reflect the latest exam Up to Date Resources for COVID 19 Exam Disruption Access to a robust online platform Comprehensive overview of the AP Chemistry exam format Hundreds of practice exercises with thorough answer explanations Proven strategies specific to each section of the test A self guided study plan including flashcards games and more online **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change Basic Chemistry William S. Seese, Guido H. Daub, 1985 *General Chemistry* Raymond E. Davis, Kenneth W. Whitten, 1996 Are you looking for the key to success in your chemistry class In CHEMISTRY you will find a strong molecular reasoning focus problem solving exercises and an innovative online homework management system that will prepare you for any challenge you might encounter The textbook is filled with learning aids that will help you master concepts of the course **Instructor's Manual for General Chemistry, Experiment and Theory** Richard H. Eastman, 1970 **5 Steps to a 5: AP Chemistry 2022 Elite Student Edition** Mary Millhollon, Richard H. Langley, 2021-08-04 MATCHES THE LATEST EXAM Let us supplement your AP classroom experience

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