

Name:

Block:

### Solutions, Acids, and Bases Test Review

#### Solutions

- 1- In soda pop,  $\text{CO}_2(g)$  is a **solute** and water is the **solvent**.
- 2- What term is used by chemists to quantitatively describe a solution in which a relatively small amount of solute is dissolved? **Dilute/unsaturated**
- 3- When a solvent has dissolved all the solute it can at a particular temperature, it is said to be **Saturated**.
- 4- What will increase the rate of dissolution of a solute in a solvent? **Stirring, increasing surface area, increasing temperature**
- 5- What is the mass of  $\text{H}_2\text{SO}_4$  in  $1.00 \times 10^2$  mL of  $0.200\text{ M}$   $\text{H}_2\text{SO}_4$  solution? **1.96g**
- 6- What number of moles of solute are present in 25.0 mL of  $2.00\text{ M}$   $\text{HCl}$ ? **0.05 Moles**
- 7- How many grams of  $\text{K}_2\text{SO}_4$  (molar mass = 174 g/mol) would be needed to prepare 4.00 L of a  $0.0510\text{ M}$  solution? **35.5g**
- 8- A solution is prepared by dissolving 7.31 g of  $\text{Na}_2\text{SO}_4$  in enough water to make 225 mL of solution. Calculate the solution molarity. **0.229M**
- 9- A 54.8 g sample of  $\text{SrCl}_2$  is dissolved in 112.5 mL of solution. Calculate the molarity of this solution. **3.07M**

#### Gas Laws

- 1- A balloon has a volume of 1.20 L at  $24.0^\circ\text{C}$ . The balloon is heated to  $48.0^\circ\text{C}$ . Calculate the new volume of the balloon. **1.340L**
- 2- A sample of helium gas occupies 2.65 L at 1.20 atm. What pressure would this sample of gas exert in a 1.50 L container at the same temperature? **2.12atm**
- 3- A gas has a volume of 5.0 L at a certain pressure. How must the pressure be changed to double the volume of the gas at constant temperature? **Pressure is 1/2**
- 4- A 6.5 L sample of nitrogen at  $25^\circ\text{C}$  and 1.5 atm is allowed to expand to 13.0 L. The temperature remains constant. What is the final pressure? **0.75atm**
- 5- You transfer a sample of a gas at  $17^\circ\text{C}$  from a volume of 5.67 L and 1.10 atm to a container at  $37^\circ\text{C}$  that has a pressure of 1.10 atm. What is the new volume of the gas? **6.06L**
- 6- A sample of helium gas occupies 12.4 L at  $23^\circ\text{C}$  and 0.956 atm. What volume will it occupy at  $40^\circ\text{C}$  and 0.956 atm? **13.1L**
- 7- A sample of oxygen gas ( $\text{O}_2$ ) has a volume of 7.51 L at a temperature of  $19^\circ\text{C}$  and a pressure of 1.38 atm. Calculate the moles of  $\text{O}_2$  molecules present in this gas sample. **0.433 moles**
- 8- A weather balloon contains helium gas at a volume of 855 L. Initially, the balloon is at sea level, where the temperature is  $25^\circ\text{C}$  and the barometric pressure is 0.947 atm. The balloon then rises to an altitude of 6000 ft, where the pressure is 0.796 atm and the temperature is  $15^\circ\text{C}$ . What is the change in the volume of the balloon as it ascends from sea level to 6000 ft? **128L**
- 9- A 6.35 L sample of carbon monoxide is collected at  $55^\circ\text{C}$  and 0.892 atm. What volume will the gas occupy at 1.05 atm and  $20^\circ\text{C}$ ? **4.82L**

# Answers To Solution Acid And Bases Test 2015

**H Kauffman**

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