

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T = temperature
 P = pressure
 V = volume
 S = entropy
 H = enthalpy
 G = free energy
 R = molar gas constant
 n = number of moles
 M = molar

m = molal
 L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E) $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M .

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United States



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Table of Contents 1999 Ap Gov Released Multiple Choice Exam

1. Understanding the eBook 1999 Ap Gov Released Multiple Choice Exam
 - The Rise of Digital Reading 1999 Ap Gov Released Multiple Choice Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying 1999 Ap Gov Released Multiple Choice Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a 1999 Ap Gov Released Multiple Choice Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1999 Ap Gov Released Multiple Choice Exam
 - Personalized Recommendations
 - 1999 Ap Gov Released Multiple Choice Exam User Reviews and Ratings
 - 1999 Ap Gov Released Multiple Choice Exam and Bestseller Lists

5. Accessing 1999 Ap Gov Released Multiple Choice Exam Free and Paid eBooks
 - 1999 Ap Gov Released Multiple Choice Exam Public Domain eBooks
 - 1999 Ap Gov Released Multiple Choice Exam eBook Subscription Services
 - 1999 Ap Gov Released Multiple Choice Exam Budget-Friendly Options
6. Navigating 1999 Ap Gov Released Multiple Choice Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - 1999 Ap Gov Released Multiple Choice Exam Compatibility with Devices
 - 1999 Ap Gov Released Multiple Choice Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1999 Ap Gov Released Multiple Choice Exam
 - Highlighting and Note-Taking 1999 Ap Gov Released Multiple Choice Exam
 - Interactive Elements 1999 Ap Gov Released Multiple Choice Exam
8. Staying Engaged with 1999 Ap Gov Released Multiple Choice Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1999 Ap Gov Released Multiple Choice Exam
9. Balancing eBooks and Physical Books 1999 Ap Gov Released Multiple Choice Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1999 Ap Gov Released Multiple Choice Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 1999 Ap Gov Released Multiple Choice Exam
 - Setting Reading Goals 1999 Ap Gov Released Multiple Choice Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 1999 Ap Gov Released Multiple Choice Exam
 - Fact-Checking eBook Content of 1999 Ap Gov Released Multiple Choice Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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