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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

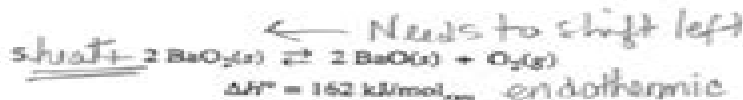
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$ (B) Removing a small amount of $BaO(s)$ X (C) Adding He gas to the vessel X (D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

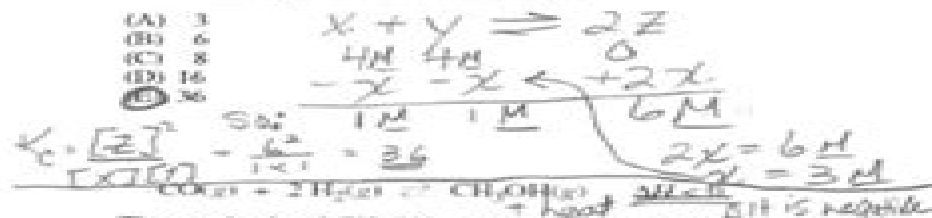
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

2002 Ap Chemistry Multiple Choice Answer Key

L Cohen



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Biology, 1999 **Science** John Michels (Journalist), 2008 A weekly record of scientific progress Heparanase Israel Vlodavsky, Ralph D. Sanderson, Neta Ilan, 2020-04-09 Written by internationally recognized leaders in Heparanase biology the book's eight chapters offer an opportunity for scientists clinicians and advanced students in cell biology tumor biology and oncology to obtain a comprehensive understanding of Heparanase's multifaceted activities in cancer inflammation diabetes and other diseases as well as its related clinical applications Proteases and their involvement in cancer progression have been well addressed and documented however the emerging premise presented within this book is that Heparanase is a master regulator of aggressive cancer phenotypes and crosstalk with the tumor microenvironment This endoglycosidase contributes to tumor mediated remodeling of the extracellular matrix and cell surfaces augmenting the bioavailability of pro tumorigenic and pro inflammatory growth factors and cytokines that are bound to Heparan sulfate Compelling evidence ties Heparanase with all steps of tumor progression including tumor initiation growth angiogenesis metastasis and chemoresistance supporting the notion that Heparanase is an important contributor to the poor outcome of cancer patients and a validated target for therapy Unlike Heparanase heparanase 2 a close homolog of Heparanase lacks enzymatic activity inhibits Heparanase and regulates selected genes that promote normal differentiation and tumor suppression Written by internationally recognized leaders in Heparanase biology this volume presents a comprehensive understanding of Heparanase's multifaceted activities in cancer inflammation diabetes and other diseases as well as its related clinical applications to scientists clinicians and advanced students in cell biology tumor biology and oncology The Encyclopedia of Mass Spectrometry Michael L. Gross, Richard Caprioli, 2006-12-21 Volume 6 Ionization Methods captures the story of molecular ionization and its phenomenal evolution that makes mass spectrometry the powerful method it is today Chapters 1 and 2 cover fundamentals and various issues that are common to all ionization e g accurate mass isotope clusters and derivatization Chapters 3-9 acknowledge that some ionization methods are appropriate for gas phase molecules and others for molecules that are in the solid or liquid states Chapters 3-6 cover gas phase molecules dividing the subject into 1 ionization of gas phase molecules by particles e g EI 2 ionization by photons 3 ionization by ion molecule and molecule molecule reactions e g APCI and DART and ionization in Strong electric fields i e Electrohydrodynamic and Field Ionization Desorption Ionization in a Strong Electric Field illustrates the transition to ionization of molecules in the solid or liquid states covered in Chapters 7-9 1 spray methods for ionization e g electrospray 2 desorption ionization by particle bombardment e g

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